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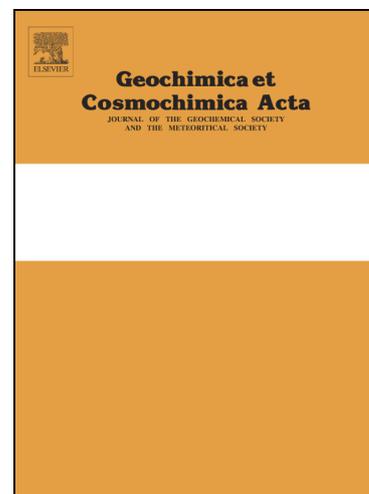
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Zinc isotopic fractionation between aqueous fluids and silicate magmas: an experimental study

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Abstract

Fluids in magmas play a key role in magma differentiation and transportation of economic metals for ore deposits. As a chalcophile and incompatible element, zinc and its isotopes have been increasingly applied to study the magmatic-hydrothermal processes. However, zinc isotopic fractionation between aqueous fluids and magmas has not been well constrained. Here we experimentally determined equilibrium fractionation factors of Zn isotopes between aqueous fluids and silicate magmas ($\Delta^{66}\text{Zn}_{\text{fluid-magma}} = \delta^{66}\text{Zn}_{\text{fluid}} - \delta^{66}\text{Zn}_{\text{magma}}$). The results reveal that aqueous fluids are isotopically heavier than the coexisting silicate magmas. No correlation between $\Delta^{66}\text{Zn}_{\text{fluid-magma}}$ and temperature or chlorine contents in fluids is observed under our experimental conditions. Instead, $\Delta^{66}\text{Zn}_{\text{fluid-magma}}$ is negatively corresponded with NBO/T of the melt (the ratio of non-bridge oxygen and tetrahedron ions), and positively correlated with the molar ratio of $\text{Al}/(0.5\text{K}+\text{Ca}+\text{Fe})$ in the bulk magmas, suggesting the controlling of silicate composition on Zn isotope fractionation. Our data therefore indicate that the isotopically heavier Zn in the fluids exsolved from magmas may account for the higher $\delta^{66}\text{Zn}$ of pegmatites and high-silica granitic rocks. Moreover, involvement of magmatic fluids can explain the highly variable and remarkably heavier Zn isotopic signatures of fumaroles, thermal spring waters, and seafloor hydrothermal fluids compared to the igneous rocks. This study provides information that can be used to guide research using Zn isotopes to trace fluid activity and magmatism.

1. INTRODUCTION

The emitted gases or aerosols by active volcanoes suggest that high temperature fluids scavenge and transport significant amounts of volatiles and metals from magma (e.g., Zelenski *et al.*, 2013; Menard *et al.*, 2014; Zelenski *et al.*, 2014). These fluids could transport fundamental components (e.g., sulfur and metals) from magma chambers to form magmatic hydrothermal ore deposits (e.g., Hedenquist and Lowenstern, 1994; Williams-Jones and Heinrich, 2005). Zinc is a strongly chalcophile and moderately incompatible element with five stable isotopes, i.e., ^{64}Zn (49.19%), ^{66}Zn (27.79%), ^{67}Zn (4.04%), ^{68}Zn (18.39%), and ^{70}Zn (0.60%). Zinc and its isotopes can be tracers in the magmatic hydrothermal systems because it is widely present in igneous rocks and hydrothermal fluids, such as fumarolic gases, condensates, and thermal springs. Zinc is a mono-valent (+2) metal, and thus its isotopic behavior is not directly affected by redox reactions (Albarède, 2004).

Zinc is a trace element in the continental crust and mantle (crust: ~72 ppm; mantle: ~55 ppm; McDonough and Sun, 1995; Rudnick and Gao, 2003). Previous studies suggested that magmatic differentiation only generates limited Zn isotope fractionation in volcanic rocks (e.g., Chen *et al.*, 2013; Huang *et al.*, 2018). In contrast, large Zn variations ($\delta^{66}\text{Zn}$ values range from 0.12‰ to 0.88‰) have been reported in felsic rocks (e.g., Telus *et al.*, 2012; Huang *et al.*, 2018; Xu *et al.*, 2019). Proterozoic to Cenozoic pegmatites studied by Telus *et al.* (2012) display some of the heaviest Zn isotope composition among all the published data of igneous rocks, with $\delta^{66}\text{Zn}$ values of +0.53‰ to +0.88‰. These Zn isotopic variations in the high-silica granite and pegmatite may be associated with the aqueous fluids exsolved from felsic magmas, but Zn isotope fractionation between fluids and magma remains unclear.

Previous studies also demonstrated that Zn isotope geochemistry has the potential to explore hydrothermal ore genesis processes (e.g., Kelley et al., 2009; Pašava et al., 2014) and trace Zn transport in hydrothermal systems (e.g., John et al., 2008; Gagnevin et al., 2012; Chen et al., 2014). Recent studies of sphalerite showed that kinetic fractionation may result in precipitation of low $\delta^{66}\text{Zn}$ sphalerite in the early fluid evolutions (Wilkinson *et al.*, 2005; Kelley *et al.*, 2009). A study on submarine hydrothermal vents also suggested that subsurface Zn sulfide precipitation prefers light Zn isotopes and can cause variations in $\delta^{66}\text{Zn}$ of hydrothermal fluids (John *et al.*, 2008). Chen et al. (2014) suggested that heavier Zn isotope favors the fluid phase during high temperature water-rock reactions. Fujii et al. (2011, 2014) and Ducher et al. (2016, 2018) calculated the fractionation factors of Zn compounds in aqueous fluids. However, the causes for large variations in Zn isotopic compositions of fumaroles from volcanic studies (Toutain *et al.*, 2008; Chiaradia *et al.*, 2018) and sphalerite in hydrothermal deposits (Mathur and Wang, 2019) are still controversial.

Equilibrium isotope fractionation factors of Zn between aqueous fluids and silicate magmas are essential for understanding Zn isotopic behaviors in magma degassing and water-rock reactions. Here we experimentally determined the Zn isotope equilibrium fractionation factors between fluids and mafic to felsic magmas (i.e., basalt, andesite, dacite, rhyolite dacite, rhyolite and haplogranite) at 700 to 900 °C and 2 kbar. These data provide Zn isotope fractionation factors between silicate magma and aqueous fluids to understand Zn isotopic behaviors in magmatic hydrothermal processes.

2. METHOD

2.1. Starting materials

Six glasses or rock powders were used as starting materials, and their compositions are given in Table 1. Basalt is a grinded natural rock from Two Buttes, southeastern Colorado, which was also used in Guo and Audetat (2017). Andesite is the natural andesite (AGV-1) and rhyolite is the natural rhyolite (RGM-1) reference materials from the USGS. Dacite and rhyolite dacite are synthesized silicate glasses with compositions similar to the volcanic rocks (Masotta *et al.*, 2016). The other felsic material is a haplogranite of 1 kbar eutectic melt composition.

2.2. High temperature and pressure experiments

All the experiments were conducted in Au capsules with 5.0 mm O.D. (outer diameter), 4.6 mm I.D. (inner diameter) and 20 mm length. Each glass or rock powder was loaded into capsules with a similar amount of solution (~100 mg) containing ~1000 ppm Zn as ZnCl₂ dissolved in water and 1.75 wt% to 7 wt% HCl. Capsules were then welded closed with a pulsed arc welder and annealed at 125 °C in an oven for a few hours, and finally weighed to check for the loss of weight. The capsules that showed no obvious weight loss (<0.0001 g) were then loaded into vertical rapid-quench cold-seal pressure vessels made of Inconel 713LC superalloy in a setup similar to that described in Matthews *et al.* (2003), using water as the pressure medium. Temperatures were measured with NiCr-Ni (K-type) thermocouples in an external borehole of the vessels. Uncertainties in the temperature and pressure readings are ± 5°C and ± 30 bar (or better), respectively. Oxygen fugacity was not specifically controlled in the vessels, but by the reaction of water with the autoclave material, it was 0.5-1 log unit above Ni-NiO buffer.

All the experiments were run at 700 to 900 °C and 2 kbar with durations ranging from 7 days to 15 days. The samples were quenched by dropping the external magnet to make the sample fall into the water-cooled zone within a few seconds.

The recovered capsules were weighted again to check the potential leaks during the experiments. The capsules without clear weight changes were then well cleaned, cooled by liquid-N₂ and punctured with a steel needle. The solution was withdrawn as much as possible with a micropipette. The capsules were opened, and boiled in deionized water for 30 minutes. After that, the samples were rinsed several times with distilled water. All of the solutions obtained during these operations were added to the solution withdrawn from the capsule. This treatment is similar to Keppler and Wyllie (1991) which is to re-dissolve material precipitated from the fluid during quenching. The residue glasses and rocks were ground and boiled in deionized water for 30 minutes to remove the fluid bubbles inside. The powders were then leached by deionized water a few times and dried in ovens for next steps. The solution samples collected in the whole process were evaporated to dryness on a hot plate for next analyses.

2.3. Analytical methods

Backscattered electron (BSE) images of polished sections through the recovered solid samples were taken on a Zeiss Gemini 1530 field emission scanning electron microscope (FE-SEM) at Bayerisches Geoinstitut, University of Bayreuth. The running conditions were under an accelerating voltage of 20 kV, a working distance of ~14 mm and an aperture of 60 μm. An energy-dispersive spectrometer (EDS) was used to identify mineral phases in the run products.

Both the sample powders and dry fluid residues were dissolved in a combination

of double-distilled, concentrated HF–HNO₃–HCl in pre-cleaned screw-top Teflon beaker. After fully digestion and evaporation of the acid solution, the residues were refluxed with 1 mL concentrated HCl. A small fraction of each sample solution was diluted in 2% HNO₃ at appropriate levels for element-composition analyses using a Perkin-Elmer ELAN DCR-II inductively coupled plasma source mass spectrometer (ICP–MS) at the CAS Key Laboratory of Crust-Mantle Materials and Environments, University of Science and Technology of China, Hefei (USTC). The detailed procedures were described in Hou and Wang (2007).

Zinc isotopic compositions of run products were analyzed at the USTC using the established procedures of Chen et al. (2016). The chemical purification of Zn was achieved by cation exchange chromatography with Bio-Rad AG-MP-1 anion resin in a 0.5 N HNO₃ media. The whole purification procedure was repeated twice to ensure efficiently purification of Zn from interference cations. The Zn recoveries through column chemistry, based on analyses of Zn concentration in the elution collected before and after the Zn cut, were >99%. Total procedural blanks were <10 ng, negligible relative to ~2 µg Zn loaded onto columns.

Zinc isotope ratios were measured using a Thermo-Fisher Neptune Plus MC-ICP-MS at a low resolution at the USTC. The purified samples were dissolved in a 2% HNO₃ solution and introduced into the instrument using an ESI PFA microflow nebulizer with an uptake rate of 50 µL/min. Zn isotope ratios are reported in standard δ -notation in per mil relative to JMC Lyon Zn standard 3-0749 L: $\delta^X\text{Zn} = [({}^X\text{Zn}/{}^{64}\text{Zn})_{\text{sample}}/({}^X\text{Zn}/{}^{64}\text{Zn})_{\text{JMC Lyon}} - 1] \times 1000\text{‰}$, where X = 66, or 68. The long term external precisions are better than 0.05‰ (2SD) for $\delta^{66}\text{Zn}$ and 0.08‰ (2SD) for $\delta^{68}\text{Zn}$. The $\delta^{66}\text{Zn}$ of USGS reference standards obtained in this study agree well within error

with previously published values (Chen et al., 2016). The results for replicate analyses also assure the accuracy and precision of our data.

3. RESULTS

The experimental conditions and Zn isotopic results of run products are shown in Table 2. The chemical compositions of run products are summarized in Table 3. The quenched experimental products contain a quenched fluid phase (fluid inclusions or cavities), and residue silicate magmas (quenched silicate melts and/or minerals) (Fig. 1). The mineral assemblages in experimental products are clinopyroxene, plagioclase, phlogopite, magnetite, and ilmenite (Fig. 1). The bulk fluid phases and quenched solid residue phases (quenched magmas) were analyzed for elemental and isotopic compositions to represent the fluids and magmas, respectively.

For Zn isotopes, the fluid phases are enriched in heavy isotopes relative to the coexisting silicate magma phases (Fig. 2), as evidenced by the data points above the 1:1 line ($\delta^{66}\text{Zn}_{\text{fluid}} = \delta^{66}\text{Zn}_{\text{magma}}$). All the analytical data for fluid and magma products are plotted on the 1:2 line on the three-isotope diagram ($\delta^{68}\text{Zn}$ vs. $\delta^{66}\text{Zn}$). Zn isotopic fractionation values between fluids and magmas ($\Delta^{66}\text{Zn}_{\text{fluid-magma}} = \delta^{66}\text{Zn}_{\text{fluid}} - \delta^{66}\text{Zn}_{\text{magma}}$) exhibit similar values for the same magma types (i.e., basalt, andesite, dacite, rhyolite dacite, rhyolite and haplogranite), demonstrated by the parallel trends of the same magma types to the 1:1 line (Fig. 2). Zn isotope fractionation values between fluids and magmas are identical ($\Delta^{66}\text{Zn}_{\text{fluid-magma}} = 0.81 \pm 0.06$ to $0.88 \pm 0.05\%$) for the haplogranite experiments using 3.5 wt% Cl bearing fluids with durations of 7 days to 15 days (Fig. 3a). For the rhyolite experiments with 1.75 to 7 wt % Cl bearing fluids, $\Delta^{66}\text{Zn}_{\text{fluid-magma}}$ also display consistent values, i.e., $0.36 \pm 0.05\%$ to $0.40 \pm 0.04\%$ (Fig. 3b). No Zn isotope variation is observed for the haplogranite experiments (3.5 wt.% Cl

in fluid) under various experimental temperature from 700 to 900 °C, with $\Delta^{66}\text{Zn}_{\text{fluid-magma}}$ vary from $0.81 \pm 0.06\text{‰}$ to $0.88 \pm 0.05\text{‰}$ (Fig. 3c). Overall, the experimental duration, fluid chlorinity, and temperature do not clearly affect the Zn isotope equilibrium fractionation between fluids and magmas under the experimental conditions in this study. However, $\Delta^{66}\text{Zn}_{\text{fluid-magma}}$ shows a broad negative correlation with NBO/T of the bulk magmas, ranging from $0.05 \pm 0.06\text{‰}$ to $1.24 \pm 0.05\text{‰}$ (Fig. 3d). Moreover, $\Delta^{66}\text{Zn}_{\text{fluid-magma}}$ increases with increasing molar ratio of $\text{Al}/(0.5\text{K}+\text{Ca}+\text{Fe})$ in the bulk magmas from $0.05 \pm 0.06\text{‰}$ to $1.24 \pm 0.05\text{‰}$ (Fig. 3e). $\Delta^{66}\text{Zn}_{\text{fluid-magma}}$ was also plotted with other elemental ratio, such as molar ratio of $\text{Al}/(0.5\text{K}+\text{Ca}+0.5\text{Na})$ (Fig. 3f), but no clear corresponding correlation was observed (Fig. S2).

4. DISCUSSION

4.1. Controlling factors for Zn isotope equilibrium fractionation

Our experimental strategy was similar to the recent copper and barium isotope studies (Guo *et al.*, 2020a; 2020b), which determined the equilibrated isotopic fractionation factors between aqueous fluids and quenched silicate magmas. Time series experiments obtained reproducible Zn isotopic fractionation factors (Fig. 3a), demonstrating that equilibrium was attained within our experimental durations. In comparison with other studies on stable isotope fractionation (e.g., Meyer *et al.*, 2008; Guo *et al.*, 2020a), our experimental durations are long enough to reach the isotopic exchange equilibrium.

Previous studies showed chloride concentrations in the fluids have remarkable effects on the partition coefficients of Zn between fluids and magmas (Urabe, 1987;

Zajacz *et al.*, 2008), i.e., $D_{\text{Zn_fluid-magma}} = C_{\text{fluid}}/C_{\text{magma}}$, where C_{fluid} is the Zn concentration in the fluid phase and C_{magma} is that in magma phase. Our results display similar trends on $D_{\text{Zn_fluid-magma}}$ increasing with increase of chloride concentration in the fluids (Fig. 4a). Increasing temperature can decrease the $D_{\text{Zn_fluid-magma}}$ in the haplogranite and 3.5 wt% Cl bearing fluids at 700-900 °C (Fig. 4b), consistent with results in a previous study (London *et al.*, 1988). However, our experiments reveal that Zn isotope fractionation is not clearly related to temperature and fluid compositions but controlled by melt compositions (Fig. 3).

The consistent $\Delta^{66}\text{Zn}_{\text{fluid-magma}}$ with different salinity fluids at various temperatures reflect the similar Zn speciation in the fluid phase. Zn-Cl complexes of ZnCl_4^{2-} and ZnCl_3^- may be stable in chloride-rich hydrothermal fluids at high temperature (e.g., Mei *et al.*, 2015; Brugger *et al.*, 2016). The overall positive $\Delta^{66}\text{Zn}_{\text{fluid-magma}}$ indicate that heavy Zn isotopes are preferentially enriched into the Zn-Cl complexes in the fluids, and light Zn isotopes are preferentially fractionated into the silicate magmas (melts and/or minerals). $\Delta^{66}\text{Zn}_{\text{fluid-magma}}$ decreases with increasing NBO/T in the melts, and increases with increasing $\text{Al}/(0.5\text{K}+\text{Ca}+\text{Fe})$ in magma components (Fig. 3e), indicating that the coordination environments of zinc in the silicates are the key factors for controlling its isotopic fractionation. As a trace element, Zn may substitute the K, Ca, and Fe cations because they are in the same period of the periodic table, and have comparable radiuses (Shannon, 1976). The silicate magmas with low NBO/T and high $\text{Al}/(0.5\text{K}+\text{Ca}+\text{Fe})$ molar ratios favor light Zn isotopes incorporating into their structures, resulting in increased $\Delta^{66}\text{Zn}_{\text{fluid-magma}}$ values, and showing the sensitive effect of the compositions of silicate magmas. Although such observations are important and useful for understanding Zn isotope fractionation mechanism, the physics behind such observations are not clear and requires further investigation.

4.2. Implications for Zn isotope compositions in igneous rocks

The felsic igneous rocks, especially high-silica granitic rocks ($\text{SiO}_2 > 70\%$), exhibit large Zn isotopic variations (Fig. 5). These high silica granites, leucogranites, or pegmatites are remarkable enriched in heavy Zn isotopes relative to the mafic igneous rocks (e.g., Telus *et al.*, 2012; Doucet *et al.*, 2018; Xu *et al.*, 2019; Wang *et al.*, 2020). The high $\delta^{66}\text{Zn}$ signature of the high silica rocks has been ascribed to the isotope fractionation during partial melting, magma differentiation and/or fluid exsolution (e.g., Telus *et al.*, 2012; Xu *et al.*, 2019; Wang *et al.*, 2020). Mantle peridotites and komatiites have indistinguishable Zn isotopic compositions (Wang *et al.*, 2017; Huang *et al.*, 2018; Sossi *et al.*, 2018), suggesting that partial melting only induces small Zn isotope fractionation, as shown by the variations from the ultramafic to mafic igneous rocks from both tholeiitic, intraplate systems and calc-alkaline systems (Fig. 5). Chen *et al.* (2013) found that Zn isotope fractionation during magma differentiation is $< 0.1\%$ based on the study of cogenetic igneous rocks from the Kilauea Iki lava lake, Hawaii and Hekla, Iceland. Rocks from the Kamchatka-Aleutian arcs (Huang *et al.*, 2018) and bimodal volcanic rocks from Hailar Basin (Xia *et al.*, 2017) only show slight Zn isotopic fractionation in magmatic systems. The magnitude of Zn isotope fractionation during fractional crystallization and partial melting are too small to explain the remarkable Zn isotope variations in high-silica granitic rocks.

Theoretical calculations have been conducted to estimate the Zn isotope fractionation between aqueous Zn species and solid Zn phase ($\Delta^{66}\text{Zn}_{\text{fluid-solid}}$) (Fujii *et al.*, 2014; Ducher *et al.*, 2016; 2018). These studies implied that the fluid phase is enriched in Zn (Zn-Cl phase) and has lighter Zn isotope composition relative to zincite (ZnO). This mineral is considered to be an analogue to the Zn coordination environment

in melt. Thus, the high-silica granitic rocks with heavier Zn isotope compositions were previously attributed to fluid removal of isotopically light Zn from the high silica magma and resulting in the heavy Zn compositions to the residue silicate rocks (e.g., Xu et al., 2019). However, our experiments reveal that fluids favor heavy Zn isotopes over the coexisting silicate magmas, with $\Delta^{66}\text{Zn}_{\text{fluid-magma}}$ as high as $\sim 0.4\text{‰}$ to $\sim 0.8\text{‰}$ for high silica rocks (Fig. 3). The discrepancy between our experimentally determined $\Delta^{66}\text{Zn}_{\text{fluid-magma}}$ and that speculated from theoretical calculation (Fujii *et al.*, 2014; Ducher *et al.*, 2016; 2018) may be due to that the Zn is differently coordinated in zincite and silicate melt. In addition, our experimental results indicate that the heavy Zn isotope composition observed in high silica rocks cannot be explained by the fluid removal process. On the contrary, the fluid formed during exsolution processes is more likely enriched in heavy Zn isotopes compared to the residue silicates. Accumulation and transportation of these fluids and reaction with wall rock afterwards might cause the heavy Zn isotope compositions of the high silica rocks.

The exsolved magma commonly experiences volatile (mainly water) saturation as it ascends and cools (Edmonds and Woods, 2018). Volatile-rich fluids are then exsolved from the magma and sequester fluid mobile elements from the magma phase to the fluid phase. These processes often happen in the highly evolved silicate magmas (high silica), as evidenced by the coexisting fluid and melt inclusions in the quartz crystals (Audétat and Pettke, 2003; Zajacz *et al.*, 2008). The aqueous fluids that exsolved from magmatic intrusions have relatively low chlorinity (2–10 wt.% $\text{NaCl}_{\text{equiv}}$) compared to the high-density brine (e.g., Audétat and Pettke, 2003; Audétat *et al.*, 2008). Natural aqueous fluids from melt exsolution is enriched Zn by forming Zn-chlorine complexes, such as ZnCl_4^{2-} and ZnCl_3^- (e.g., Mei *et al.*, 2015; Brugger *et al.*, 2016). Assuming that the fluid phase is continuously removed from the magma through fractional distillation, then

$\delta^{66}\text{Zn}$ of residue magmas and the fluids during this process can be modelled by the following Rayleigh fractionation equations:

$$\delta^{66}\text{Zn}_{\text{magma}} = (\delta^{66}\text{Zn}_{\text{initial}} + 1000) f^{(\alpha - 1)} - 1000 \quad (1)$$

$$\delta^{66}\text{Zn}_{\text{fluid}} = (\delta^{66}\text{Zn}_{\text{initial}} + 1000)(f^\alpha - 1)/(f - 1) - 1000 \quad (2)$$

Where $\delta^{66}\text{Zn}_{\text{initial}}$ refers to the Zn isotopic composition of the initial volatile-rich magmas and f is the fraction of Zn in residual magmas, α is the Zn isotope fractionation factor between the fluid and magma, with the relationship $\alpha \approx 1 + \Delta^{66}\text{Zn}_{\text{fluid-magma}}$. The model is conducted with $\Delta^{66}\text{Zn}_{\text{fluid-magma}} = 0.2, 0.4, \text{ and } 0.8\%$, which are the ranges of the experimental results. The model shows that fluid exsolution (commonly from very low content to up to 10 wt.%) can produce the fluids with similarly heavy Zn isotopic compositions compared to the high silica granitic rocks (Fig. 5).

Magmatic volatiles physically separate from the parent magmas and form discrete liquid pulses that both form new fractures, and follow existing, fractures into the brittle rocks overlying the batholithic system (Tosdal and Richards, 2001). The narrow conduits could be an accumulation zone for the volatiles and metals rich fluids, because the dense magma forms a downward return flow to the large magma batholith. Fluid pressure can further cause fractures in the cover rocks, and fault zones provide high-permeable pathways for fluid and magma ascending. The fractures may first be opened and propagated by the fluid pressure, and later filled by more viscous magma, forming cylinder dikes of high silica granite or pegmatite (Burnham, 1997; Carrigan et al., 1992; Rubin, 1995). The felsic rocks formed in the volcanic conduits are likely to inherit the high $\delta^{66}\text{Zn}$ features of magmatic fluids, because the environment is enriched in fluids containing substantial abundance of Zn. Pegmatites are usually formed at the roofs of

the intrusions where aqueous fluids coexist with silicate melts, which is evidenced by the common occurrence of miarolitic cavities (e.g., Thomas et al., 2009; London et al., 2012). The high silica rocks are altered by the magmatic fluids enriched in heavy Zn isotopes, and should inherit the heavy Zn isotopic compositions of the fluids.

4.3. Implications for Zn isotopes in hydrothermal systems

Zinc is a moderately volatile element that behaves similarly to the volatiles (e.g., S and Cl) and transition metals (e.g., As, Sb, and Tl) during magma degassing (Symonds *et al.*, 1987; Taran *et al.*, 1995; Guo and Audétat, 2017). Thus, magmatic fluids, which contain considerable amounts of Zn, should be the dominant Zn source for the fumarole gases. Zinc isotopic variations in the fumarolic gases may be due to the isotopic fractionation during the degassing processes. Since magma differentiation (segregation of crystals in the magma chambers) induces limited Zn fractionation, the Zn isotopic signatures of the magmas with fumarolic components are assumed to be homogeneous. The degree of Zn incorporation into the fumarolic gases is defined in terms of relative enrichment factors (EF) (Lepel *et al.*, 1978):

$$EF = (X_g/R_g) / (X_m/R_m) \quad (3)$$

For an element X, g and m refer to the gas sample and the rock, respectively, and R is a fluid-immobile reference element with respect to volcanic fluids. Since Ti was highly immobile in the volcanic fluids, Ti was usually chosen as the reference element. The enrichment factors of Zn range from 120 to 1041 obtained from the analyses of high temperature fumarolic gases/aerosols and igneous rocks from arc volcanoes such as Tolbachik (Zelenski et al., 2014), Erta Ale (Zelenski et al., 2013), Lascar (Menard et al., 2014) and Etna (Aiuppa *et al.*, 2003). The high Zn concentration in fumarolic

components suggests that magma degassing could enrich Zn in the fumarolic gases, which could also fractionate Zn isotopes between the magmas and the fumarole gases.

The fumarolic gases (Toutain *et al.*, 2008; Chen *et al.*, 2014; Chiaradia *et al.*, 2018), thermal spring waters of La Soufrière volcano (Chen *et al.*, 2014), hydrothermal alteration zones around the crater fumaroles (Vulcano) (Chiaradia *et al.*, 2018), and sea floor hydrothermal vents (John *et al.*, 2008) show large variations but generally high $\delta^{66}\text{Zn}$ compared to the mafic to intermediate silicate rocks (Fig. 5). This is broadly consistent with our experimental results that magma-derived fluids preferentially scavenge heavy Zn isotopes from magmas. The large variations and low $\delta^{66}\text{Zn}$ values are probably due to the water rock reaction, mixing with water from other sources, and/or kinetic fractionation during sphalerite precipitation (Wilkinson *et al.*, 2005; Chen *et al.*, 2014; Chiaradia *et al.*, 2018).

Toutain *et al.* (2008) firstly reported that $\delta^{66}\text{Zn}$ in the fumarolic gases of Merapi volcano (Indonesia) are ranging from +0.05‰ to +0.85‰, generally higher than the $\delta^{66}\text{Zn}$ of homogeneous andesitic rock samples with a mean value of +0.24‰. The condensate samples from the gases display very high $\delta^{66}\text{Zn}$ from +1.48‰ to +1.68‰. They proposed that the condensates are enriched in the heavier isotopes from the gasses and thus condensation decreases gaseous Zn concentrations enriched in light Zn isotopes. Chen *et al.* (2014) also reported that fumaroles from La Soufrière volcano, on Guadeloupe Island (French West Indies) have similar or slightly higher $\delta^{66}\text{Zn}$ (+0.21‰ to +0.35‰) compared to the local bedrock fresh andesite (+0.21‰), and thermal springs are enriched in heavy Zn isotopes compared to the host rocks (Fig. 5). Although Zn stable isotope fractionation may also happen during the hydrothermal reactions at

high temperatures, these studies imply that the initial fluids derived from magma degassing have higher $\delta^{66}\text{Zn}$ than their host magma reservoirs.

The Zn isotope compositions ($\delta^{66}\text{Zn}$ ranges from -0.3‰ to +0.2‰) of Vulcano fumaroles are overall lighter than these of fumaroles from Merapi and La Soufrière (Chiaradia *et al.*, 2018). A mixture of a hydrothermal endmember with low $\delta^{66}\text{Zn}$ and a magmatic endmember with high $\delta^{66}\text{Zn}$ are suggested to account for the trend (Chiaradia *et al.*, 2018). Chiaradia *et al.* (2018) also found that the hydrothermal alteration zones around the fumarole vents have significantly high $\delta^{66}\text{Zn}$ (0.3–1.6‰), probably due to that the initial fumaroles are enriched in heavy Zn isotopes. Our experimental results are consistent with the volcanic observations that magma-derived fluids or gases are characterized by high $\delta^{66}\text{Zn}$ than their host magma reservoirs. Zn isotopic fractionation during magma degassing could be a novel tool to constrain the hydrothermal processes in the magmatic differentiation, volcanic eruption, and element precipitation during water-rock reactions.

Except the volcanic systems, both high and low $\delta^{66}\text{Zn}$ have been discovered in hydrothermal deposits (Gagnevin *et al.*, 2012; Mathur and Wang, 2019). These anomalous Zn isotope ratios have been widely explained by precipitation of isotopically light sphalerite (Zn sulfide) in the early stage during kinetic fractionation from hydrothermal fluids (Wilkinson *et al.*, 2005; John *et al.*, 2008; Kelley *et al.*, 2009). However, the heavy $\delta^{66}\text{Zn}$ fluids derived from magmas will become even heavier in Zn isotopes after the Zn isotopically light sulfide precipitation, which will be recorded by the subsequent sulfides at various stages. The large Zn isotope fractionation may be used to detect new deposits (Gagnevin *et al.*, 2012), trace hydrothermal fluid sources (Pašava *et al.*, 2014), or to identify the different mineral deposit stages (Wilkinson *et*

al., 2005; Mathur and Wang, 2019). The experimental results and the observed high $\delta^{66}\text{Zn}$ in the hydrothermal systems suggest that the magma generated fluids could be an important source of high $\delta^{66}\text{Zn}$, which could contribute to global Zn cycles through interactions with groundwater, rivers, and atmosphere. The contribution of hydrothermal systems should be taken account when studying the global oceanic budget of Zn using Zn isotope data.

5. CONCLUSIONS

This study provides the first experimental study on Zn isotope equilibrium fractionation between aqueous fluids and silicate magmas at magmatic-hydrothermal conditions. In this study, chloride bearing fluids and silicates were equilibrated at conditions relevant to the coexistence of hydrothermal fluids and silicate magmas. Experiments involving fluids of different salinity, temperature, and rock composition resulted in a wide range of $\Delta^{66}\text{Zn}_{\text{fluid-magma}}$ from $0.05 \pm 0.06\%$ up to $1.24 \pm 0.05\%$. The magnitude of Zn isotopic fractionation between fluid and magma is primarily depended on the magma compositions. While Zn species in aqueous fluids remain constant within a normal range of salinity and P-T conditions, the fractionation may be controlled by the Zn substitution mechanism in the melt structure.

Our data can be used to interpret the evaluated $\delta^{66}\text{Zn}$ values of high-silica granitic rocks that involve the Zn in magmatic fluids. The Zn isotopic composition detected in samples derived from hydrothermal systems are broadly consistent with our experimental results. The heavy Zn isotopic signatures of magma-derived fluids can be observed in the fumarole samples, thermal springs or hydrothermal vents. Zn isotopes in hot hydrothermal fluids can be also affected by the rapid precipitation of sulfide segregation in isotopically light Zn isotopes. By deciphering the geochemical processes

that affect Zn isotopic behaviors in the hydrothermal systems, our study provides new data that can be applied to Zn isotope studies of fluid relevant geological processes.

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Figure captions

Fig. 1. Back-scattered electron (BSE) images of run products in the experiments: (a) basalt from run Zn17 after a run time of 7 days at 850 °C and 2 kbar. (b) andesite from run Zn7 after a run time of 7 days at 850 °C and 2 kbar. (c) dacite from run Zn16 after a run time of 7 days at 850 °C and 2 kbar. (d) rhyolite dacite from run Zn19 after a run time of 7 days at 850 °C and 2 kbar. (e) rhyolite from run Zn12 after a run time of 10 days at 850 °C and 2 kbar. (f) haplogranite from run Zn12 after a run time of 13 days at 800 °C and 2 kbar. Abbreviations: cpx - clinopyroxene; plg - plagioclase; phl - phlogopite; mag - magnetite; spl - spinel.

Fig. 2. The $\delta^{66}\text{Zn}$ values in experimental magma products vs. in fluid products at 700 to 900 °C and 2 kbar with durations ranging from 7 days to 15 days. The results show the recovered fluids have higher $\delta^{66}\text{Zn}$ values than the residue magmas, and their fractionation factors are quite comparable for the same starting materials.

Fig. 3. Isotopic fractionation values ($\Delta^{66}\text{Zn}_{\text{fluid-magma}}$) of Zn between fluid and magma as a function of (a) experimental durations from 7 to 15 days in a haplogranite, (b) salinity in the starting fluid (1.75-7.0 wt% Cl) in a rhyolite, (c) temperature from 700 to 900 °C in a haplogranite, (d) NBO/T of the starting materials, (e) rock compositions ($\text{Al}/(0.5\text{K}+\text{Ca}+\text{Fe})$ in mole), and (f) rock compositions ($\text{Al}/(0.5\text{Na}+0.5\text{K}+\text{Ca})$ in mole) at all conditions.

Fig. 4. Partition coefficient (D_{Zn} fluid-magma) of Zn between fluid and magma as a function of (a) salinity in the starting fluid (1.75-7.0 wt% Cl) in a rhyolite, (b) temperature from 700 to 900 °C in a haplogranite, (c) NBO/T of the starting materials, and (d) rock composition ($\text{Al}/(0.5\text{K}+\text{Ca}+\text{Fe})$ in mole) at all conditions.

Fig. 5. (a) Rayleigh fractionation model of $\delta^{66}\text{Zn}$ in fluids and magmas during magma degassing. A fractional degassing model was used and the initial $\delta^{66}\text{Zn}$ values in the silicate magmas was assumed to be $0.20 \pm 0.05\text{‰}$ (around the average Zn isotopic composition of Bulk Silicate Earth, according to Wang et al. (2017) and Sossi et al. (2018)). $\Delta^{66}\text{Zn}_{\text{fluid-magma}}=0.2, 0.4$ and 0.8 ‰ were used in the modelling. (b) A summary of Zn isotopic compositions for the igneous rocks and geological samples relevant to the magmatic fluids. Zn isotopic compositions of the igneous rocks are from Telus et al. (2012), Chen et al. (2013), Doucet et al. (2018), Xia et al. (2017), Huang et al. (2018) and Xu et al. (2019); Other samples' Zn isotopes are from John et al. (2008), Toutain et al. (2008), Chen et al. (2014) and Chiaradia et al. (2018).

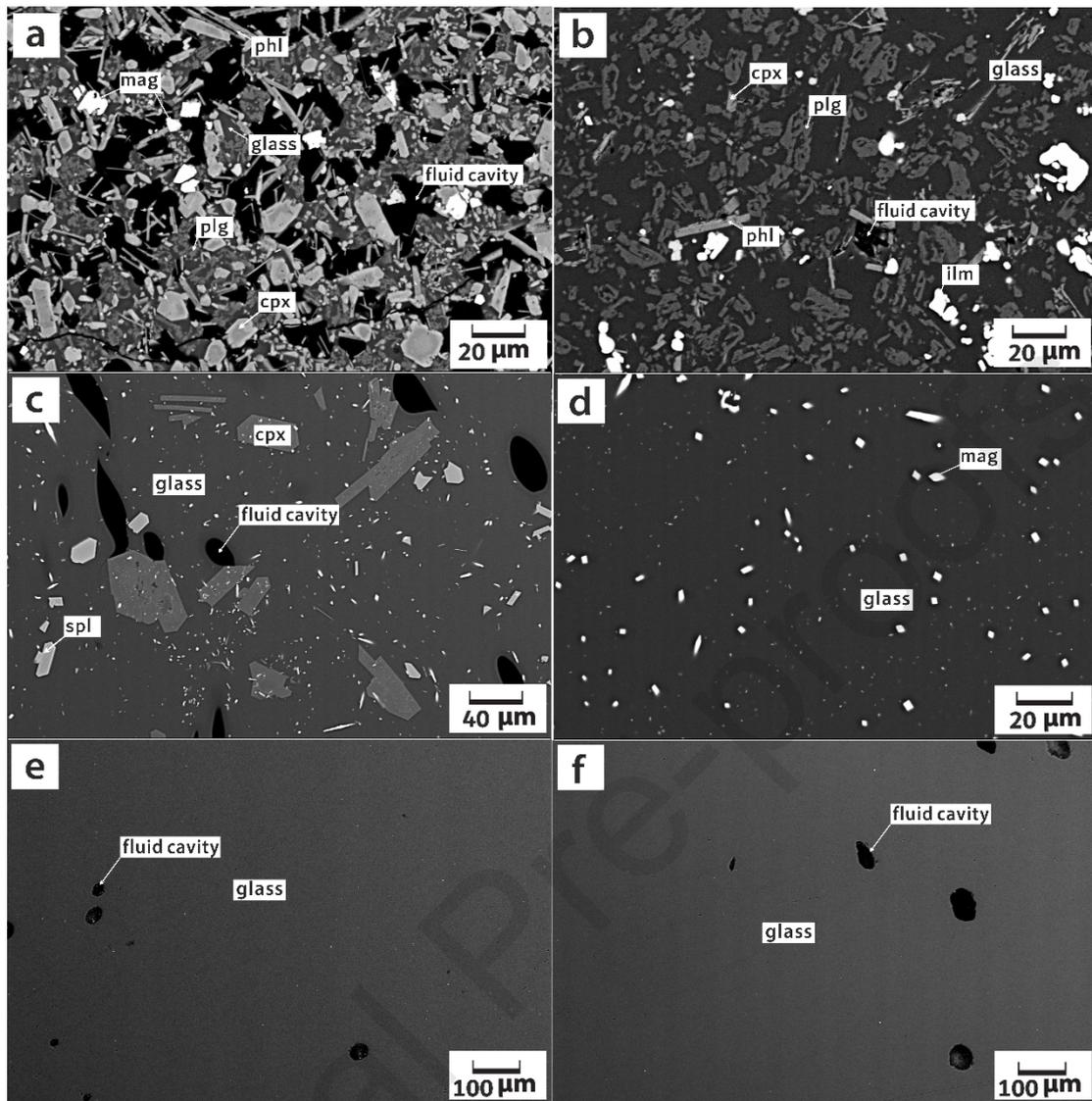


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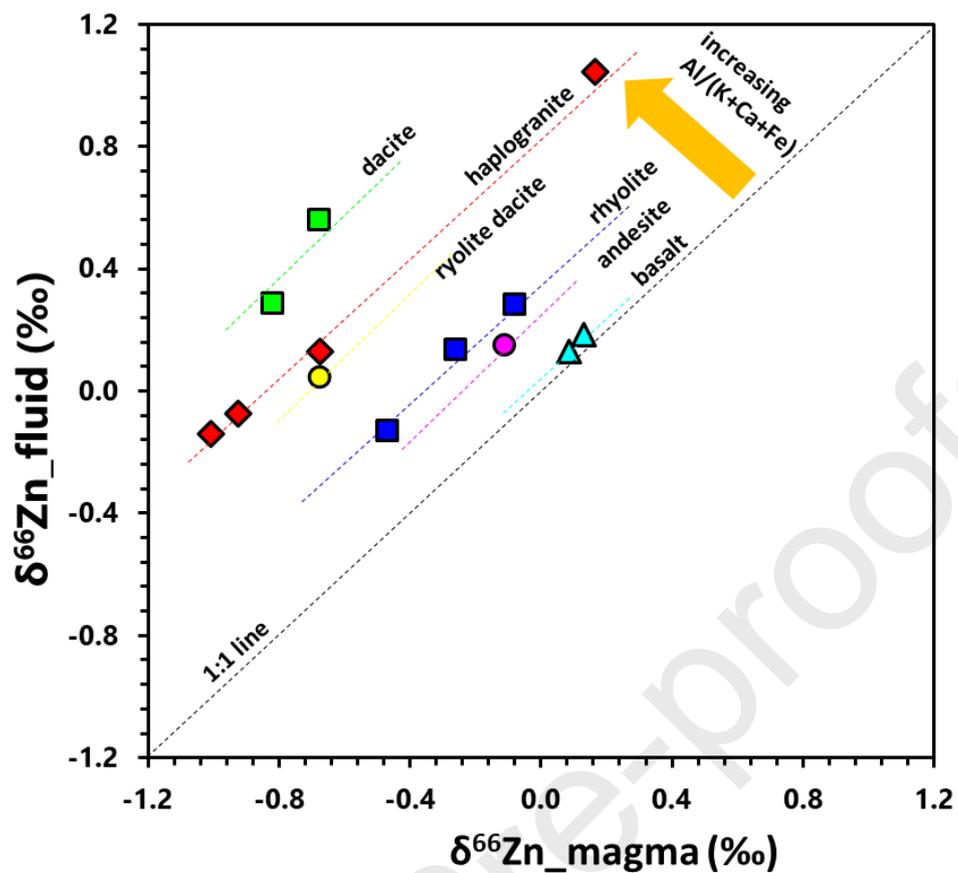


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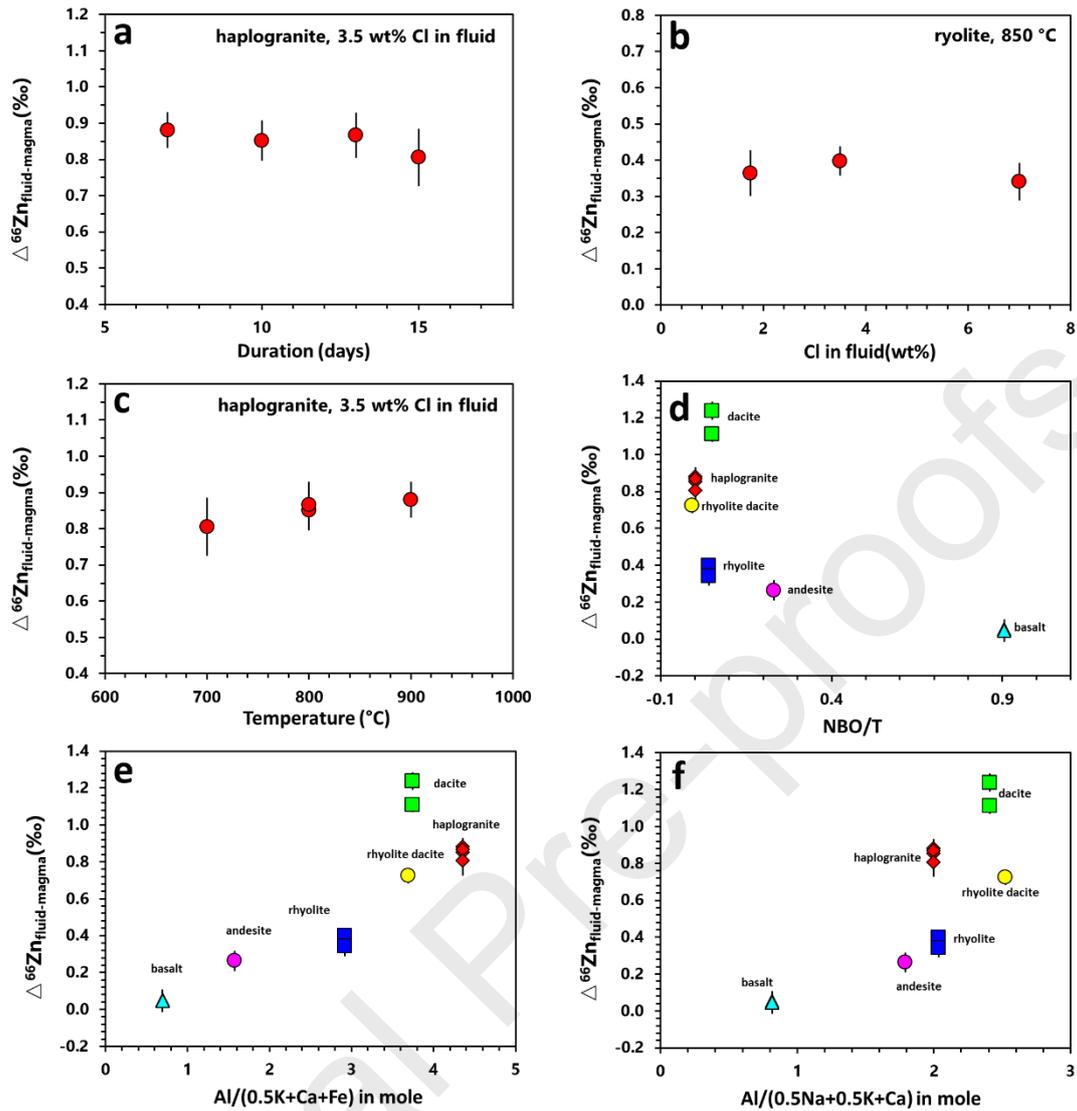


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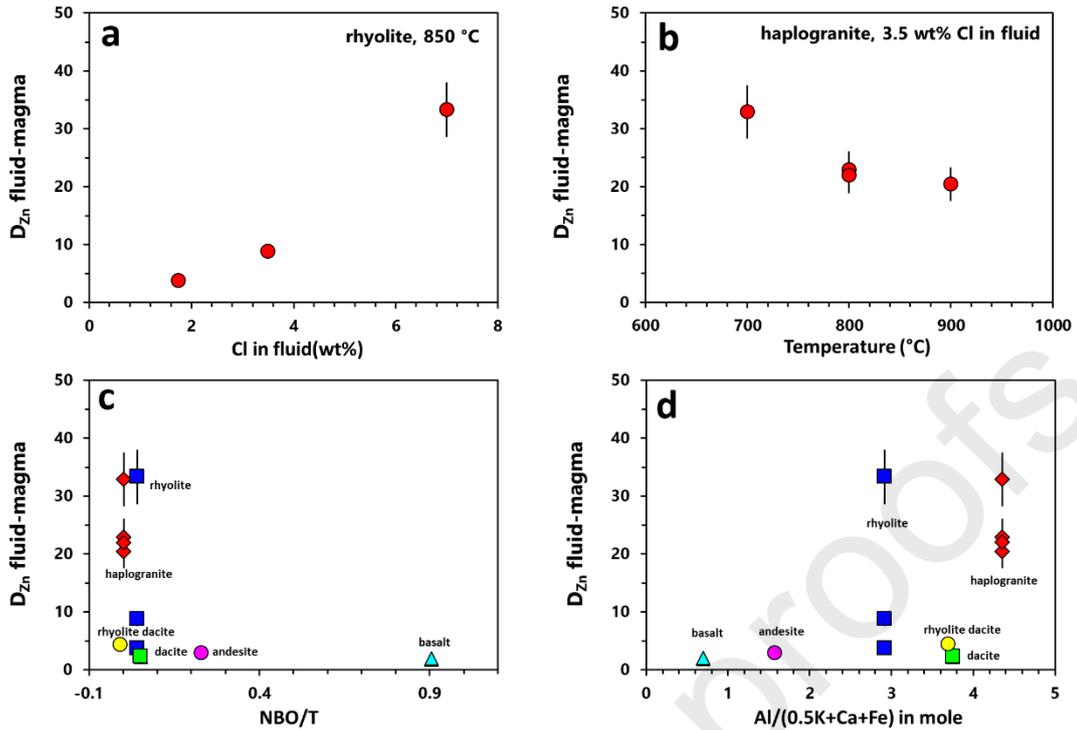


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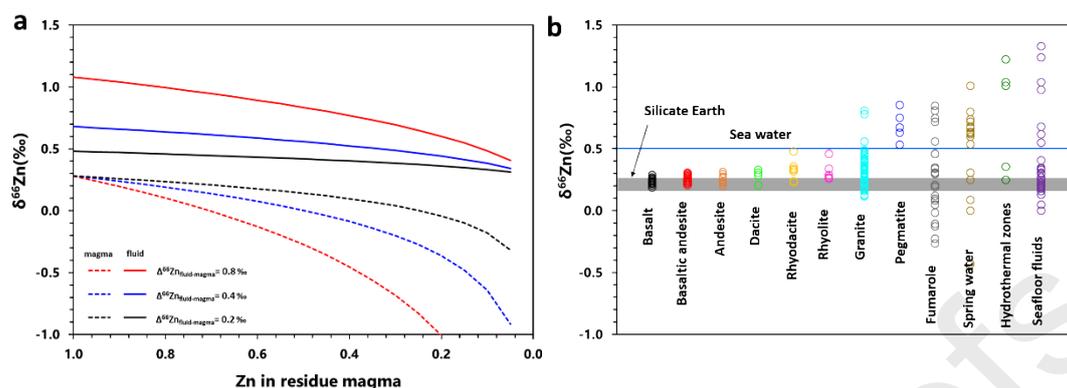


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Table 1

Compositions of starting materials. All oxides are in weight percent.

Element	basalt	andesite	dacite	ryolite dacite	ryolite	haplogranite
Na ₂ O	2.3	4.26	5.4	3.6	4.1	3.9
MgO	10.6	1.53	1.5	0.3	0.28	0.02
Al ₂ O ₃	11.1	17.15	18.8	16.1	13.7	11.7
SiO ₂	54.1	58.8	66.0	72.2	73.4	79.5
K ₂ O	4.1	2.92	2.1	3.4	4.3	4.9
CaO	10.4	4.94	2.5	1.8	1.15	
TiO ₂	1.0	1.05	0.7	0.6	0.27	0.02
FeO _t	6.1	6.77	2.3	1.3	1.86	0.04

Table 2

Experimental conditions and results overview.

run No.	T (°C)	P (kbar)	duration (days)	starting material	Cl in solution (wt%)	$\delta^{66}\text{Zn}_{\text{magma}}$	2SD	$\delta^{68}\text{Zn}_{\text{magma}}$
Zn04	850	2	7	ryolite	1.75	-0.08	0.04	-0.15
Zn12	850	2	10	ryolite	3.5	-0.26	0.02	-0.52
Zn05	850	2	7	ryolite	7	-0.47	0.02	-0.90
Zn08	900	2	7	haplogranite	3.5	0.17	0.01	0.31
Zn09	800	2	10	haplogranite	3.5	-0.92	0.02	-1.79
Zn13	800	2	13	haplogranite	3.5	-1.01	0.04	-2.00
Zn10	700	2	15	haplogranite	3.5	-0.67	0.03	-1.30
Zn17	850	2	7	basalt	3.5	0.13	0.01	0.22
Zn11	850	2	7	basalt	3.5	0.08	0.03	0.19
Zn07	850	2	7	andesite	3.5	-0.11	0.03	-0.20
Zn16	850	2	7	dacite	3.5	-0.82	0.03	-1.55
Zn16.1	850	2	10	dacite	3.5	-0.68	0.04	-1.29
Zn19	850	2	7	ryolite dacite	3.5	-0.68	0.01	-1.15

Table 3

Recovery solid phase (quenched magma) and fluid compositions analyzed by ICP-MS.

Sample	Zn4		Zn12		Zn5		Zn8		Zn9	
	magma	fluid								
Li	59	11.2	52	11.7	42	23.9	2	3.4	1.9	1.3
Na	25,310	4,080	21,689	5,362	17,654	9,852	21,824	919	20,758	2,387
Mg	1578	15	1502	27	1374	100	98	n.a.	92	n.a.
Al	71,680	179	66,071	256	66,891	434	59,602	175	63,340	270

K	31,445	4,146	26,418	6,288	22,034	11,413	32,705	592	31,093	2,696
Ca	2,970	36	2,681	44	2,607	159	156	12	125	2
Sc	4.6	n.a.	4.4	n.a.	4.1	n.a.	1.0	n.a.	1.0	n.a.
Ti	1,009	11	1,256	7.3	901	8.4	133	1.4	135	1.0
V	3.9	n.a.	3.2	n.a.	3.1	n.a.	n.a.	n.a.	n.a.	n.a.
Cr	2.0	n.a.	n.a.	n.a.	n.a.	n.a.	n.a.	n.a.	0.65	n.a.
Mn	254	70	173	107	82	202	4	1	3	1
Fe	10,668	2,075	7,318	4,718	3,326	8,471	216	41	397	136
Co	6.9	5.6	7.2	11	3.9	15	92	6	3	3
Ni	95	48	148	167	83	530	398	79	246	326
Zn	175	661	98	864	29	967	45	920	26	892
Rb	139	20	113	31	97	57	2	0.5	2	0.3
Sr	120	11	110	12	106	16	125	11	129	11
Ba	887	81	824	67	804	111	36	74	36	46

n.a.- not analyzed

Declaration of competing interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

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