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Microfluidic flow-through reactor and 3D Raman imaging for *in situ* assessment of mineral reactivity in porous and fractured porous media[†]Jenna Poonosamy,^{*a} Cyprien Soulaire,^b Alina Burmeister,^c Guido Deissmann,^a Dirk Bosbach^a and Sophie Roman^bReceived Date
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An in-depth understanding of dissolution and precipitation of minerals in porous and fractured porous media and the complex feedback on the transport of fluids is essential for various subsurface applications. In this context, we developed a novel non-destructive “lab-on-chip” approach for quantitative *in situ* assessments of mineralogical changes in porous media. Our experimental approach involves a microfluidic flow-through reactor of reactive homogeneous and heterogeneous (fractured) porous media coupled with high-resolution imaging. Here, the reactive medium consists of compacted celestine grains seeded in a reservoir within the microfluidic chip. This medium reacts with a barium chloride solution injected into the microreactor at a constant flow rate, leading to the dissolution of celestine and growth of barite. Various seeding processes of the mineral grains allow the creation of homogeneous reactive porous media or the introduction of large heterogeneities such as fractures. Hence, our approach enables high-resolution investigations of reactive transport in fractured porous media. The use of confocal Raman spectroscopic techniques enables the spatio-temporal visualization of the mineral transformation at the pore-scale in two- and three-dimensions. Moreover, advanced pore-scale modelling correlates the hydrological heterogeneities to the geochemical observations in the micro-reactor, which explains the observed discrepancies between homogeneous and heterogeneous reactive media. Eventually, the proposed methodology can be applied to other chemical systems to provide new insights into hydro-geochemical coupling in porous and fractured porous media as well as high-fidelity datasets to benchmark reactive transport codes that are currently under development.

Introduction

Reactive transport modelling^{1,2}, is a versatile tool in predicting or assessing mineralogical changes due to reactive fluids in space and time for various subsurface applications including shale gas production using hydraulic fracturing (“fracking”), CO₂ sequestration, geothermal energy extraction, the underground repositories for nuclear waste, or environmental remediation^{3,4}. Because of its widespread applications, recent studies have been dedicated to test the underlying concepts of these numerical models using experimental benchmarks^{5–9}. Discrepancies between nu-

merical models and experimental observations were noted and mainly explained by the fact that reactive transport operates at the continuum-scale (averaged properties of the porous media, e.g. porosity, permeability, effective surface area) and considers upscaled parameters^{10,11}, while pore-scale features (e.g. surface roughness of minerals¹²) influence the macroscopic behaviour (e.g. transport properties) of the system. These experiments suggest the need to understand the processes occurring at the pore-scale and to develop process-based predictive models and mathematical relationships that account for small-scale heterogeneities¹³.

In the last decades, there has been a growing interest to use pore-scale modelling in combination with experimental imaging/characterisation of porous media as tools to get a better understanding of coupled chemical and transport processes and to upscale the results to a macroscopic representation^{2,14–17}. A live monitoring of chemical processes and hydraulic pathways can be performed using non invasive techniques such as magnetic res-

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onance imaging¹⁸, positron emission tomography¹⁹, or neutron imaging²⁰. The spatial resolution, however, is often a limiting factor for these techniques. Alternatively, high resolution experimental images can be used. For example, Prasianakis *et al.*²¹ investigated the effect of pore size distribution on the nucleation of sulphates in porous media using scanning electron microscopy (SEM) images of the reacted porous media. Noiriél *et al.*²² and Molins *et al.*²³ combined pore-scale modelling with X-ray microtomography imaging to assess calcite dissolution. These imaging techniques may nevertheless require specific facilities (e.g. Synchrotron) or can also be destructive without providing any information on the pore architecture or the spatial distribution of minerals of the original porous micro-structure (e.g. post-experimental characterization using SEM).

Another option that allows a direct visualization of flows, reactions and transport mechanisms at the pore-scale involves microfluidic devices coupled with optical microscopy techniques. Micromodels are two-dimensional representations of a porous medium in which the pore space is etched on a substrate by photolithography methods and then covered by a transparent material. The etched pattern may be of variable complexity ranging from single micro-channels of different shapes to patterns representative of natural porous media obtained from imaging of real rocks²⁴. For now, microfluidic platforms are mainly based on glass, silicon or PDMS (a silicone elastomer) materials. However, transport and reaction dynamics depend also on rock surface chemistry. Recently, micromodels that include reacting minerals have emerged. They enable direct visualization of the fundamental fluid-mineral interactions at the pore-scale. For example, Song *et al.*²⁵ etched flow channels in a natural calcite crystal substrate that was bonded to a glass plate to investigate carbonate dissolution by an acidic brine. Soullaine *et al.*¹⁵ investigated the complex feedback between hydrodynamics and surface chemistry using a single calcite crystal posted in a microchannel. Singh *et al.*²⁶ developed a real-rock/PDMS combined micromodel allowing them to investigate transport properties and multiphase flow in real-time.

In this paper we propose a new kind of micromodels based on a flow-through setup. It consists of a microfluidic reservoir filled with reacting crystals of controlled grain size and mineralogy with a reacting fluid flowing through it. The reservoir can contain either a homogeneous or a heterogeneous (fractured) porous medium. Our monitoring techniques include optical microscopy and Raman spectroscopy and allow in situ, non destructive and real-time monitoring of details on the chemical composition of the system at the grain-scale. The aim of this study is to assess the use of a new generation of microfluidic flow-through reactors in understanding mineral reactivity in porous and fractured porous media.

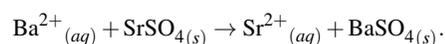
Materials and methods

In this section we describe our experimental setup and the measurement techniques that we developed to investigate mineral reactivity during micromodel experiments.

Experimental concept

The experimental setup consists of a micro-reactor connected to a pump and monitored by optical microscopy (Fig. 1). We chose a well-controlled chemical system¹³, which is neither pH nor redox sensitive. The reactor includes a 'crystal reservoir' filled with celestine (SrSO₄) crystals. A barium chloride (BaCl₂) solution is injected in the reactor through the inlet at a constant flow rate. The injection of the BaCl₂ solution triggers the dissolution of celestine and subsequently barite (BaSO₄) precipitation due to the difference in their solubility products ($K_{sp}BaSO_4 = 10^{-9.97} \text{ mol}^2\text{L}^{-2}$ and $K_{sp}SrSO_4 = 10^{-6.63} \text{ mol}^2\text{L}^{-2}$ at 298.15 K and 1013 hPa, from Hummel *et al.*²⁷).

The chemical reaction writes as follows,



Micromodel

The microfluidic reactor consists of a reservoir of 900 μm length and 800 μm width in which the reactive crystals are injected. The reservoir contains five supporting square pillars of 50 μm length to enhance the mechanical stability. The reservoir is placed in between two sections (upstream and downstream, see Fig. 1) that constitute an artificial pore architecture made of cylindrical pillars of 100 μm diameter and pore throats of 1 μm. The pore throat of 1 μm enables the confinement of the crystals in the reservoir during the seeding process (see next section). Moreover, the upstream section (in the absence of heterogeneities in the distribution of crystals in the reservoir) enables an uniform flow of the solution as it reaches the reactive zone and the downstream section prevents that precipitates plug the outlet. The upstream and downstream sections are connected to the inlet and outlet respectively via channels of 2600 μm length and 100 μm width. The reservoir is connected to a crystal injection channel of 1900 μm length and 300 μm width. The crystal injection channel was made wide enough such that crystals do not get trapped but can be flushed freely during the seeding process. The depth of the microfluidic device is 10 μm, close to the size of the crystal which allows them to get stacked in the reservoir. The silicon wafer, produced by photolithography²⁸, contains the positive relief of the micromodel and is used as a mold in the PDMS (Polydimethylsiloxane), Silicone Elastomer, Sylgrad) molding process. A cast of the micromodel is obtained by pouring transparent liquid PDMS onto the mold, baking it (1 h, 65 °C), and removing it from the wafer. Holes for the fluids interconnections are manually punched in the PDMS. Before bonding to the glass plate, the PDMS layer is cleaned with isopropanol, dried under nitrogen flow and treated with adhesive tape²⁹. An oxygen plasma treatment is applied to the PDMS and glass layers in order to activate their surface. Then, they are assembled and heated for 5 minutes at 80 °C until they are securely bonded.

Celestine seeding process

The celestine crystals used in the experiments are obtained from naturally occurring celestine SrSO₄ stones of 99.7% purity⁷. The stones were crushed and sieved to collect crystals of < 20 μm grain

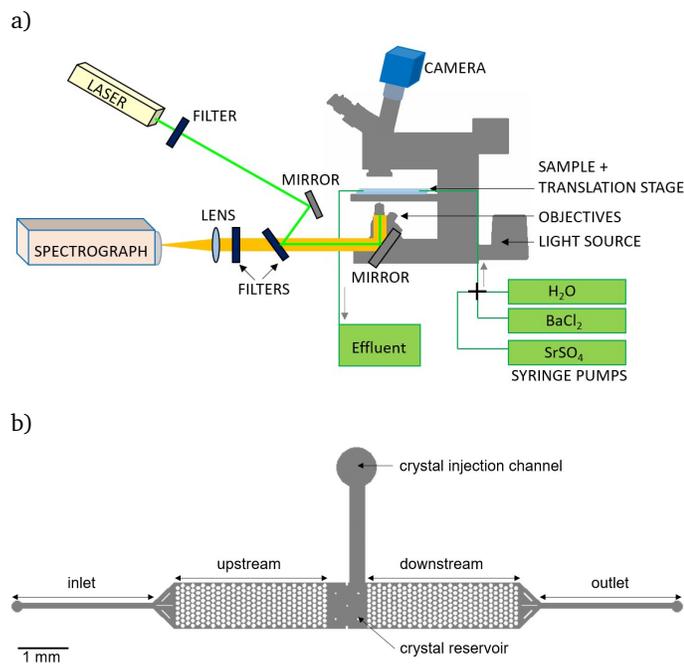


Fig. 1 a) Schematic of the experimental setup b) geometry of the micro-model, inlet and outlet channels are 0.10 mm wide and 2.50 mm long, upstream and downstream porous media are 0.90 mm wide and 3 mm long, the crystal reservoir is 0.90 mm wide and 0.80 mm long, the crystal injection channel is 0.3 mm wide and 1.9 mm long, the depth is 10 μm everywhere.

size. This fraction of crystals was further segregated and crystals of grain size between 4 μm and 9 μm were isolated using sedimentation techniques (in isopropanol) in an Atterberg cylinder³⁰. The celestine crystals were mixed in a saturated solution of strontium sulphate (SrSO_4) to constitute a seeding suspension. Prior to the injection of the seeding suspension, the microfluidic device was flushed with water to prevent gas bubbles being trapped in the system. The injection of the crystals using this technique can be used to generate a homogeneous compacted porous medium in the reservoir (Fig. 2a). By alternating the injection of the seeding suspension and the injection of a SrSO_4 solution through the inlet, a heterogeneous (fractured) compacted porous medium can be obtained (Fig. 2b). The injection channel was mechanically closed using a stopper (Microfluidic luer lock stopper from Darwin Microfluidics, Delauney, France). A saturated SrSO_4 solution was then injected at a constant flow rate of 1000 nLmin^{-1} for 3 hours to stabilize the crystals in the reservoir, and later at 500 nLmin^{-1} for 21 hours to ensure a steady-state condition, using a syringe pump (NeMeSYS, Cetoni GmbH, Germany). This step ensures that the crystals were well stabilized and less prone to colloidal movements in the reservoir. Two experiments were conducted, involving a homogeneous and a heterogeneous compacted porous medium, labelled A and B, respectively.

Fluidics

The inlet was connected to three dispensing syringes (water, a saturated solution of SrSO_4 and a 1 mM BaCl_2 solution) using a cross peek connector (052P-722, CS Chromatographie Ser-

vice GmbH, Langerwehe, Germany) and the outlet was linked to an effluent vessel. The micromodel was initially saturated with SrSO_4 solution (see previous section). Then, water was injected at 500 nLmin^{-1} for 30 minutes to prevent precipitation of barite in the upstream section. Finally, BaCl_2 solution was injected into the microfluidic device at 500 nLmin^{-1} for 60 hours.

Imaging and Raman spectroscopy

The microfluidic experiment was observed using an automated inverted microscope (Witec alpha300 Ri Inverted Confocal Raman Microscope which consists of an inverted Nikon Ti-2 U as base microscope) with 100 \times oil immersion objective of numerical aperture (NA) 1.25, working distance 0.23 mm and cover glass correction for 0.17 mm from Nikon. The instrument is equipped with a Nd:YAG laser ($\lambda = 532 \text{ nm}$), and a thermoelectrically cooled CCD Camera and Ultra-High-Throughput Spectrometer UHTS300. The lateral and axial (depth) resolution of Raman measurements with the focus at the sample surface is usually estimated from the theoretical diffraction limit given by $d_l \approx \frac{1.22\lambda}{NA}$ and $d_a \approx \frac{4\lambda}{NA^2}$, yielding 520 nm and 1362 nm, respectively.

Before the reaction, high resolution optical images of the compacted porous medium or the fractured medium were taken and the images were stitched for visualization of the porous media structure (see fig. 2 a and b). In each experiment, two regions of interest (e.g. compacted regions or regions along a fracture) were selected and the mineralogical changes were monitored at regular time intervals using Raman spectroscopy, see fig. 2 c and d. *In situ* hyperspectral Raman images were recorded from an area of 60 $\mu\text{m} \times 50 \mu\text{m}$ with a 260 nm step size (232 pixel \times 192 pixel per image) in x and y direction. The individual Raman spectra were collected during continuous x-y stage movement. Raman intensities were recorded for 0.1 s in the wavenumber range from 300 to 1300 cm^{-1} . During the Raman measurements, the adjustable laser power was set to 50 mW and the grating to 1800 grooves min^{-1} with a 500 nm entrance slit. The total exposure time was 1 hour and 43 minutes for a single image. The selected map area was a good compromise of the representative elementary volume (< 1% pore volume exchange) and acquisition time. At the end of the experiment, the mineral mappings of the reservoir and a depth scan along a 100 μm line in the upstream region were carried out to visualize the minerals distribution in the cell. The step size were 2000 nm (475 pixel \times 475 pixel per image) in x and y direction, and 3840 nm step size in x and 384 nm in z direction (384 pixel \times 42 pixel per image), respectively. In addition, stacked Raman images were collected 25 $\mu\text{m} \times 50 \mu\text{m}$ with a 260 nm step size (96 pixel \times 192 pixel per image) in x and y direction and over a depth of 13 μm with a 1000 nm step size. These images were used to reconstruct the 3D geometry of the porous media skeleton.

Raman image stacks were visualized with Volume Viewer, an ImageJ plugin developed by Kai Uwe Barthel Internationale Medieninformatik, HTW Berlin (Germany), that shows stacks as slices, projections, or volumes within a 3D-space. We also used ImageJ 3D Viewer plugin³¹. An image processing workflow was developed using ImageJ to measure the porosity and volume fraction of each mineral from the stacks. For that, each mineral phase was

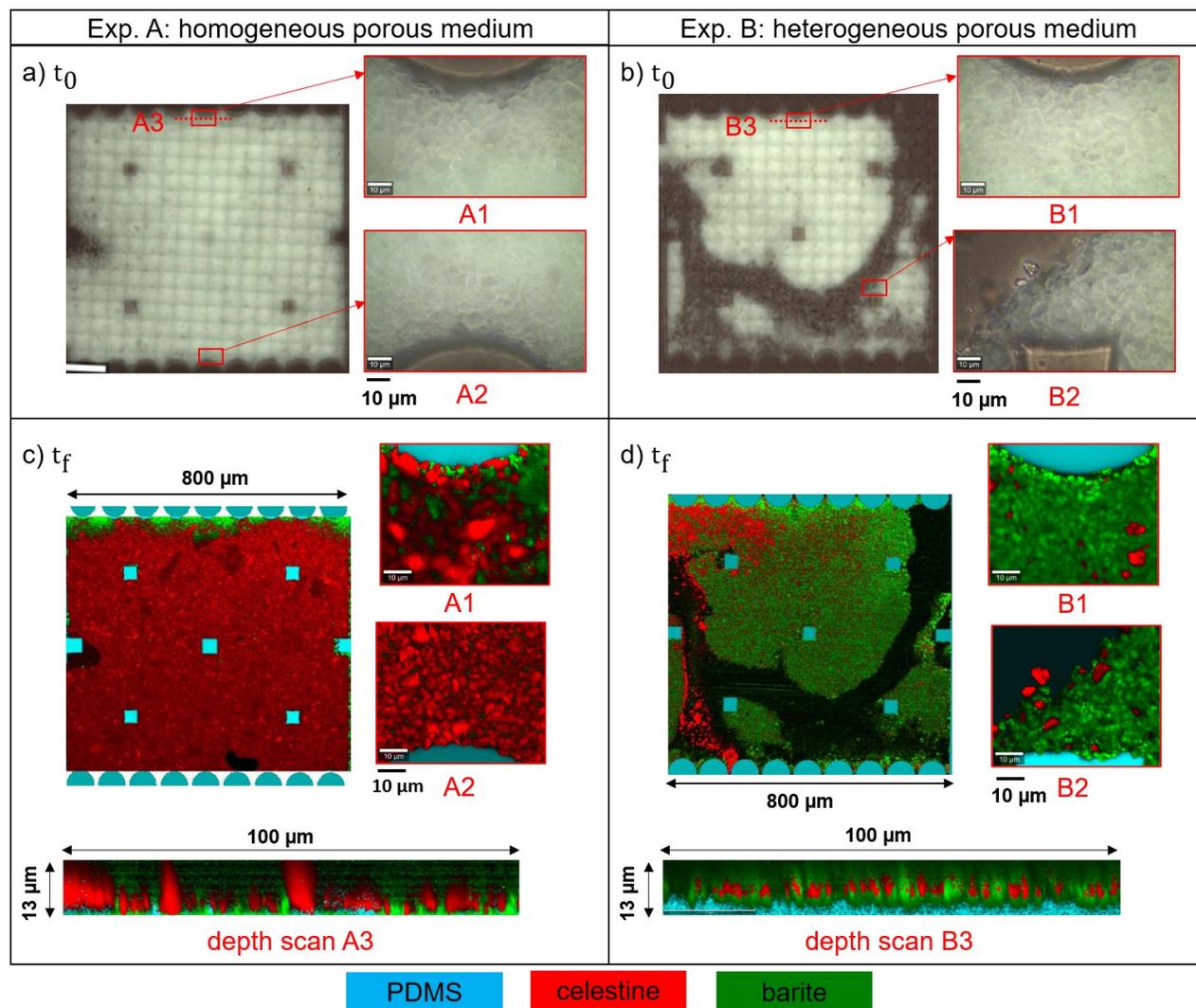


Fig. 2 a) Optical image of the reservoir filled with crystals constituting a homogeneous porous medium (experiment A) with a zoom on the monitored areas A1 and A2, b) optical image of the reservoir filled with crystals constituting a heterogeneous (fracture occupying 30% of the domain) porous medium (experiment B), with a zoom on the monitored areas B1 and B2, c) Raman imaging of experiment A with a zoom on the monitored areas A1, A2 and depth scan A3 after 60 hours, d) Raman imaging of experiment B with a zoom on the monitored areas B1, B2, and depth scan areas B3 after 60 hours. The flow direction is from top to bottom

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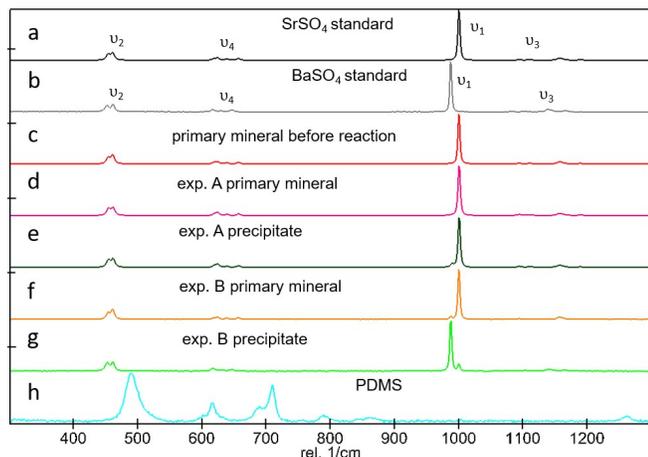


Fig. 3 Average Raman Spectra of a) SrSO_4 , b) BaSO_4 , c) unreacted celestine, d) primary celestine crystals in experiment A, e) newly formed precipitates covering primary undissolved crystals in experiment A, f) undissolved primary minerals in experiment B, g) newly formed precipitate with relics of the undissolved crystals in experiment B h) PDMS

converted to grayscale levels and segmented.

Spectra analysis of the sulphate minerals and PDMS

The Raman spectra of synthetic BaSO_4 (99.99% from Chempur), SrSO_4 (99.99% from Chempur) and the cured PDMS were collected for 0.4 s in the wavenumber range from 200 to 1400 cm^{-1} . These measurements served as standards for further evaluation of our experimental data using the True Component Analysis module from Witec Control 5.6 Software. The free sulphate ions SO_4^{2-} have a T_d symmetry and therefore nine modes of vibrations³² represented as $A_1 + E + 2F_2$ with $\nu_1(A_1)$ and $\nu_3(F_2)$ corresponding to the symmetric and the anti-symmetric stretching mode respectively, and $\nu_2(E)$ and $\nu_4(F_2)$ corresponding to the bending vibrations (Figure 3). The line positions depend on the nature of the cations, the temperature and pressure. The Raman spectra of the commercial BaSO_4 and SrSO_4 are shown in Fig. 3. Although, the spectra of BaSO_4 and SrSO_4 are rather similar, the very intense peak (ν_1) for barite and celestine are distinct at 989 and 1001 cm^{-1} , respectively, and enables their identification. Under the given experimental conditions, barite and celestine can form solid solutions^{33,34} which are usually characterised by a shift of their respective peaks, bringing them closer to each other and a broadening of their width. According to theoretical calculations (see supplemental information in Poonosamy *et al.*¹³), solid solutions can be expected under our experimental conditions. However, the main focus of this study is to test the current setup for assessing the spatial/temporal mineralogical changes of the porous media; therefore solid solution composition is not addressed. The Raman spectrum of PDMS (Fig. 3h) is different to that of the sulphate minerals and therefore does not interfere with the identification of the minerals.

Numerical evaluation of velocity fields

Numerical simulations were used to assess the local streamlines and velocity magnitudes in the microfluidic cell. This knowl-

edge allows us to relate the different reaction rates observed in the porous matrix between experiments A and B to the transport mechanisms in the micromodel. Indeed, even though the fluid and crystal properties, as well as the injection rate, are the same for both cases, the spatial distribution of the crystals in the reservoir can change drastically the distribution of the local flow rates. Because of the large contrast of characteristic length scales between the fracture and the porous matrix formed by the assembly of the celestine crystals, a full Navier-Stokes solution in the entire domain is too computationally demanding and requires the exact knowledge of the three-dimensional structure of the porous regions. Instead, we used the micro-continuum Darcy-Brinkman-Stokes approach developed by Soulaïne and Tchepeli³⁵ for which Stokes equations are solved within the fracture and the flow is described using Darcy's law in the matrix. Optical images of the microfluidic cell such as Fig. 7a are segmented into a solid phase (the white pillars), the void space that includes fractures and the spaces between the solid pillars (light grey) and the porous matrix (dark grey) and directly used to populate the micro-continuum model following the procedure introduced in Soulaïne *et al.*³⁶. The permeability and porosity of the celestine matrix are set to $2 \times 10^{-14}\text{ m}^2$ ¹³ and 0.865 (from Raman tomography measurement), respectively. In both cases, we use a constant injection velocity set to $9.25 \times 10^{-4}\text{ m/s}$ that corresponds to the mean velocity for the imposed flow rate in the experiments. The simulations are two-dimensional and the effects of the top and bottom plates of the micromodels are modelled adding a Hele-Shaw correction in the flow equations^{24,37}. The micro-continuum model is implemented and solved using OpenFOAM, an open-source computational fluid dynamics platform that uses the finite-volume method to solve partial differential equations.

Results and Discussion

In this section, we evaluate the spatio-temporal mineralogical changes in experiment A and B and use the simulated velocity field to discuss the observed changes. Finally, we discuss uncertainties and optimization potential of Raman measurements.

Mineral distribution in the system

Mineral phases are identified by the characteristic positions and relative intensities of their fundamental Raman bands (Fig. 3). The Raman signature of the celestine before reaction is shown in (Fig. 3c). Fig. 2c and d show the distribution of the remaining celestine (identified in red) and newly formed barite (identified in green) in experiments A and B. The average Raman spectra of the red regions are given in Fig. 3d (experiment A) and Fig. 3f (experiment B), while that of green regions are given in Fig. 3e (experiment A) and Fig. 3g (experiment B).

In experiment A, a $50\text{ }\mu\text{m}$ wide precipitation front was observed at the upstream side of the reservoir, leaving the majority of celestine undissolved after the reaction time of 60 hours. The precipitation of barite in experiment A, is characterised by a low intensity peak at 989 cm^{-1} , as a shoulder to the high intensity celestine ν_1 peak (Fig. 3e). Barite also precipitated sparsely downstream in between the celestine grains (Fig. 2c area A2). Based on the 3D

representation of experiment A (see supplementary material), we can infer that barite precipitated in the pore solution (homogeneous nucleation).

In experiment B, the mineral distribution is heterogeneous (Fig. 2d) with a higher barite content close to the fracture and higher content of celestine in the left side of the reservoir. The average Raman spectra of the green regions (Fig. 3g) shows the signature of barite but also relics of the undissolved celestine crystals. The undissolved celestine crystals can be easily distinguished in the depth scans, where the Raman signal from celestine is seen. An average of the red regions identified as undissolved celestine shows minor contributions of barite (Fig. 3f) suggesting barite precipitation on its surface (heterogeneous nucleation). Comparing the depth scan A3 and B3, the chemical reactions upstream (defined by celestine dissolution and consequential barite precipitation) appear more pronounced in experiment B than in experiment A (see further assessment in Fig. 6).

Fig. 4 shows the mineralogical changes in regions B1 and B2 with time. The precipitation of barite starts on the step like features of the crushed celestine crystals (Fig. 4 at 10 hours), this is a common feature observed for barite epitaxial growth on celestine³⁸ which is explained by the reduction of enthalpies due to the minimization of interfacial energies³⁹. Barite precipitates on the surface of celestine and in between the celestine grains. From the temporal changes of Area B2, it appears that the reaction occurs in the compacted regions first rather than in regions directly in contact with the fracture, where crystals (see encircled crystal in Fig. 4) dissolve, shrink and become rounder.

3D representation of the porous structure

Fig. 5 shows 3D representations of the porous micro-structure (Area A1 and B1 of experiment A and B, respectively) obtained from the stacked Raman images over a depth of $10\ \mu\text{m}$. The initial porosity was determined at 0.86 for area A1 in experiment A and 0.87 for area B1 for experiment B. After celestine dissolution and barite precipitation, the porosity increases to 0.91 for experiment A and decreases to 0.84 for experiment B. The volume fraction of minerals before and after reaction, measured for the images in Fig. 5, is shown in Fig. 6. The volume fraction of barite after the end of the experiment is 0.0017 and 0.16 in area A2 and B2, respectively, indicating a greater conversion of celestine to barite in experiment B. The replacement of celestine by barite is usually accompanied by a porosity decrease because of the larger molar volume of barite^{7,13}. This decrease in porosity is indeed observed in experiment B, while in experiment A, a porosity increase is observed suggesting a faster dissolution that precipitation in during the time of the experiment. The 3D animated representation of the porous skeletons are presented in supplementary materials.

Influence of the hydrodynamic heterogeneity on mineral reactivity

The different barite distributions in experiments A and B are attributed to the difference in velocity fields in the two systems. In experiment A, the velocity field is homogeneous with a mean velocity of $9.25 \times 10^{-4}\ \text{ms}^{-1}$ in the porous reservoir. This velocity is

calculated from the injection flow rate and cross sectional area of the micromodel. In experiment B (Fig. 7 and 2c), the velocity field is heterogeneous with the fastest flow path occurring along the fractures at $3.16 \times 10^{-3}\ \text{ms}^{-1}$ and the lowest flow rates observed in the compacted porous regions with an averaged value of $4.90 \times 10^{-5}\ \text{ms}^{-1}$. These results show that, for the same injection rates, the mean flow velocity through the celestine crystals exhibit orders of magnitude difference depending on whether the crystals are disposed uniformly in the reservoir or if they form a fractured porous medium. Indeed, we calculated that the flow rate through the compacted celestine crystals about a factor of 19 smaller in the heterogeneous case (experiment B) than in the homogeneous case (experiment A). The Peclet numbers (Pe, defined as the ratio of advective and diffusive transport rates) in the compacted regions were calculated at 4.63 and 0.25 considering a characteristic length $5\ \mu\text{m}$ (mean pore size) for experiment A and experiment B, respectively. Therefore, the transport in the porous medium is advection-dominated in experiment A and diffusion-dominated in experiment B. The observed higher volume of barite that precipitated in experiment B (in Area B1) compared to experiment A (in Area A1) can be explained by the lower velocities, lower Peclet number, thus by higher residence time that prevails in experiment B. Despite the same saturation index (logarithm of the ratio of activity of aqueous components over solubility product) in both experiments, barite precipitates in the pore space in experiment A and on the surface of the celestine (in the compacted porous regions) in experiment B. The saturation index that prevails in the system as the 1 mM BaCl_2 reaches the celestine crystals is 3.35, whereby heterogeneous nucleation should be favored over homogeneous nucleation (see Fig. 14 in Poonosamy *et al.*¹³). However, it seems that the higher flow rates in experiment A compared to the flow rates in the porous region in experiment B hinder the formation of barite clusters on the surface of celestine but add a mixing effect in the pore solution, which could favor homogeneous nucleation (mixing induced precipitation in porous media⁵).

Experiment A is also characterised by a sharp precipitation front. The rationale could be that aqueous barium in the pore solution drops to the solubility limit of barite and consequently there is not enough barium in solution for further precipitation downstream. This is a phenomena that was observed in Poonosamy *et al.*¹³. However, this is not the case here, as the number of moles of barium injected $1.8 \times 10^{-5}\ \text{mol}$ is higher than the amount of celestine $3 \times 10^{-12}\ \text{mol}$ present in the reactor. Therefore, it is most likely that the precipitation of barite downstream is kinetically hindered. Indeed as the pore solution gets depleted in aqueous barium, despite supersaturation, the energy barrier to nucleate further downstream might be too important to be overcome in the given short residence time (i.e. induction time for barite nucleation is much higher than the residence time). Homogeneous nucleation rate of barite drops abruptly to zero at saturation index of 3 (Fig. 14 in¹³). This is also in agreement with recent work on barite whereby authors reported that nucleation of barite require high supersaturation levels⁴⁰ and delayed nucleation in confined space^{21,38,41,42}.

In experiment B, the mineral transformation occurs in regions

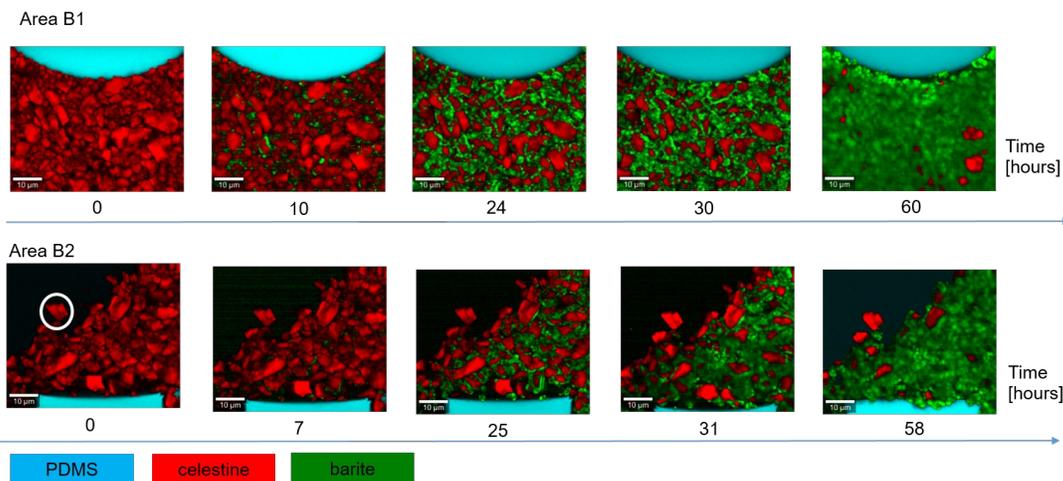


Fig. 4 The temporal evolution of the mineralogical changes of experiment B in area B1, and area B2. The white circle shows a crystal that undergoes dissolution with time

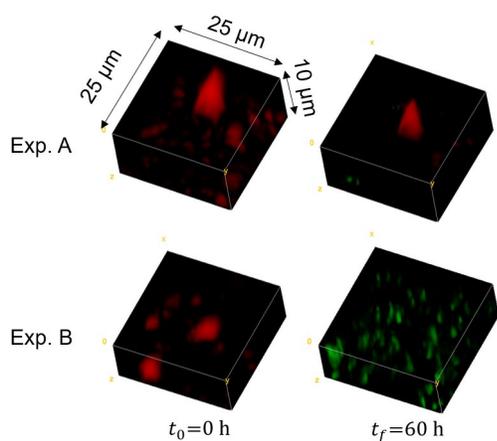


Fig. 5 The temporal evolution of the pore architecture in experiment A and B.

close to the main flow path (the fracture) with the left side of the reservoir without barite precipitation. In the regions close to the fracture, there is a continuous exchange between the pore solution and the fluid in the fracture, resulting in an increase in celestine dissolution and barite precipitation. The presence of a fracture also gives rise to local heterogeneity in the chemical reactions. This is emphasized in Fig. 4, where the temporal distribution of the minerals in area B1 and B2 is shown. A higher amount of barite precipitates on the compacted celestine crystals while the crystals along the fractures (Fig. 4 encircled crystal) seem to undergo dissolution. In the compacted regions, the flow velocities are lower increasing the residence time and therefore giving time for the concentration of Ba^{2+} and SO_4^{2-} to increase in the pore solution and barite to precipitate. The constantly renewed solution in the fracture, is undersaturated with respect to strontium sulphate, driving the dissolution of celestine. Besides, due to relatively high velocities along the fracture (short transport time compared to induction time for barite nucleation); it is unlikely that barite nucleates on the surface of celestine directly adjacent

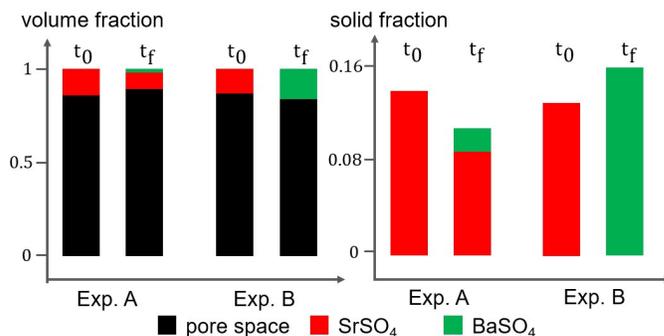


Fig. 6 evolution with time of volume fraction of celestine (red), barite (green) and pore space (black) in experiment A and B, in area A1 and B1 respectively

to the fracture. It is clear that the different transport velocities in both experiments affect the precipitation mechanisms in the porous media; and this will be further explored in future work.

Uncertainties and optimization potential of Raman measurements

In our study, a laser power of 50mW was chosen, since such a laser power provides an excellent signal to noise ratio. However, we estimated that successive mapping is likely to increase the amount of dissolved celestine by less than 3 % for the derived temporal temperature increase of about 7 C, compared to isothermal conditions (see Supplementary Material for details). To mitigate this effect for the given example (celestine to barite transformation), the laser power could be reduced to 30mW, as a clear distinction between celestine and barite can be made due to the high intensity of the ν_1 bands. In addition, the mapping time can also be decreased by considering a step size of 500nm instead of 260nm. Such optimization of the Raman measurements should be considered in the future to reduce or avoid potentially perturbing effects of laser induced heat on the experiments and the studied processes, such as increased mineral dissolution or faster

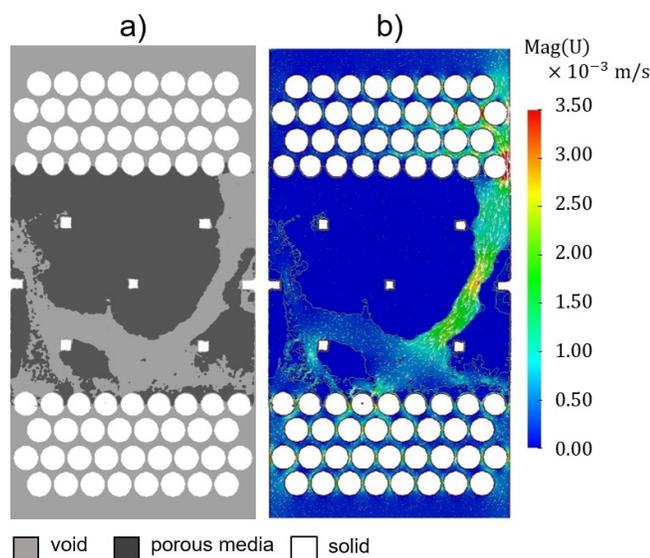


Fig. 7 a) 2D representation of the domains of experiment B implemented in the numerical simulations b) associated velocity fields computed by the micro-continuum model.

reaction kinetics.

Conclusions

We developed and established a new flow-through microfluidic reactor coupled with high-resolution imaging that enables *in situ* and non destructive 3D assessment of mineralogical and microstructural changes with full spatio-temporal resolution on the grain scale. The micro-reactor consists of compacted reactive grains seeded in a reservoir within the microfluidic chip. Various degrees of heterogeneities can be considered ranging from uniform distribution of the grains to fractured porous media. In previous studies using macroscopic flow-through reactors to investigate the impact of supersaturation on the barite overgrowth on celestine, the ongoing dissolution and precipitation processes as well as the porosity evolution were inferred from post experimental destructive investigation. In comparison, this methodology allows a systematic study of the chemical and porosity evolution of the system at the pore-scale. In comparison, this methodology allows a systematic study of the chemical and porosity evolution of the system at the pore-scale. The presented micronized flow-through reactor monitored by non-invasive imaging Raman spectroscopy is a promising tool in capturing local processes for the interpretation of coupled chemical and transport processes observed at the macro-scale. In particular, our methodology allows the investigation of reactive transport processes in fractured porous media in well-controlled environments. The current setup can be extended to other chemical systems (e.g. calcite dissolution followed by gypsum precipitation, calcite by fluorite; gypsum by barite or celestine) for rationalizing coupled chemical processes and effect of different transport regimes (in terms of Peclet numbers and/or Damköhler numbers), on the mineral reactivity in (fractured) porous media. In this work, we have shown that it was possible to micronize conventional flow-through reactors

to assess mineralogical changes at the pore scale, such concept can also be applied to micronize conventional diffusion cells such as those used, e.g. Chagneau *et al.*⁸ or Rajyaguru *et al.*⁹. Our methodological protocol is therefore a versatile tool in support of further developments in reactive transport modelling for assessing mineralogical changes and consequential porosity changes of porous media. Indeed, as reported in a recent review¹⁰, reactive transport models rely on the use of empirical relationships to predict the system evolution in terms of mineralogical changes and transport properties. The heterogeneity in porous media likely controls the mass fluxes and residence time in the reactive media and therefore are known to control the effective rates of mineral dissolution by reducing effective mineral surface areas physically (e.g. in fractures, or at partial saturation) and chemically^{43–46}. Moreover, the reactive surface area of minerals constituting the porous media is not directly experimentally accessible and is often used as a fitting parameter in reactive transport simulations⁴⁷. In a future work, the Raman images of the evolving crystals can be used to evaluate the geometry and size of the crystals with time leading to calculation of dissolution rates. Moreover, pore-scale images can be used in pore-scale modelling. In fact, our in-situ 3D Raman reconstruction of the pore architecture along with advanced pore-scale modelling¹⁴ will provide a better assessment of the transport properties (e.g. permeability, dispersion tensor) as well as the accessible reactive surface area and reaction rates for describing mineral precipitation/dissolution in porous media.

Conflicts of interest

There are no conflicts to declare.

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Contribution of authors

JP designed the micromodel. JP and SR designed and conducted the experiments and processed the experimental results. CS conducted the modelling. JP, SR and CS discussed, interpreted the

results and wrote the paper. AB fabricated the silicon wafer. GD and DB applied for funding (HGF, BMBF). All authors revised the manuscript.

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