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Evidence for mantle metasomatism by hydrous silicic melts derived from subducted oceanic crust

[Letters to Nature]

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Abstract

The low concentrations of niobium, tantalum and titanium observed in island-arc basalts are thought to result from modification of the sub-arc mantle by a metasomatic agent, deficient in these elements, that originates from within the subducted oceanic crust 1. Whether this agent is an hydrous fluid 2 or a silica-rich melt 3 has been discussed using mainly a trace-element approach 4 and related to variable thermal regimes of subduction zones 5. Melting of basalt in the absence of fluid both requires high temperatures and yields melt compositions unlike those found in most modern or Mesozoic island arcs 6,7. Thus, metasomatism by fluids has been thought to be the most common situation. Here, however, we show that the melting of basalt under both H₂O-added and low-temperature conditions can yield extremely alkali-rich silicic liquids, the alkali content of which increases with pressure. These liquids are deficient in titanium and in the elements niobium and tantalum and are virtually identical to glasses preserved in mantle xenoliths found in subduction zones 6 and to veins found in exhumed metamorphic terranes of fossil convergent zones 7. We also found that the interaction between such liquids and mantle olivine produces modal mineralogies that are identical to those observed in metasomatized Alpine-type peridotites 8. We therefore suggest that mantle metasomatism by slab-derived melt is a more common process than previously thought.

A mid-ocean ridge basalt from the Juan de Fuca ridge (Table 1) was reacted between 10 and 30 kbar with bulk H₂O contents of 2 to 10 wt% (see ref. 9 for methods), that is, largely in excess of the H₂O content of metamorphosed basalts, and at temperatures below 1,000°C. Recent experimental work has shown that melting under these conditions (addition of H₂O, low temperatures) is required to account for the geochemical and physical characteristics of slab melts 9, and approaches the most likely pressure-temperature (P-T) conditions for slab melting (750 °C, 30 kbar) as constrained from thermal modelling 10.

	P (kbar)	T (°C)	H ₂ O*	SiO ₂	TiO ₂	Al ₂ O ₃	FeO†	MgO	CaO	Na ₂ O	K ₂ O	Gl	Hbl	Cpx	Gt	Tmt	Ilm	Rut	Sph
Experimental glasses from hydrous basalt melting																			
Starting‡	–	–	–	49.9	2.3	13.7	12.7	6.8	10.6	2.5	0.3	–	–	–	–	–	–	–	–
27	10	841	9.9	69.8	0.2	18.7	0.9	0.5	5.8	3.4	0.6	41.1	26.0	21.7	–	8.6	2.6	–	–
38	20	800	9.2	68.8	0.3	16.7	2.9	1.0	5.9	3.4	0.7	28.0	67.4	–	–	–	0.1	–	7.0
39	20	900	10.2	65.4	0.4	20.9	1.7	0.2	7.0	3.7	0.6	36.2	39.1	20.9	–	–	4.0	–	–
40	20	1,000	10.0	59.3	1.5	20.1	5.1	0.9	7.5	4.3	0.8	46.8	26.5	20.8	–	5.9	–	–	–
41	30	900	6.4	70.9	0.3	16.1	0.5	0.1	1.4	9.3	1.5	23.9	–	42.9	31.8	–	0.4	0.9	–
42	30	1,000	6.0	68.0	0.6	17.9	1.2	0.2	3.7	7.1	1.1	25.6	–	34.8	38.4	–	–	1.1	–
Felsic veins in mantle xenoliths or in metamorphic complexes from subduction zones																			
Val55/4	–	–	–	67.8	0.4	18.0	0.6	0.2	1.1	7.9	2.1	–	–	–	–	–	–	–	–
Val55/4	–	–	–	63.9	0.9	19.7	2.7	1.0	3.0	5.7	0.9	–	–	–	–	–	–	–	–
41867A	–	–	–	72.2	0.1	16.5	0.2	0.1	2.8	6.6	0.3	–	–	–	–	–	–	–	–

Analyses and phase proportions given in wt%. Glass analyses normalized to 100% anhydrous. Gl, glass; Hbl, hornblende; Cpx, clinopyroxene; Gt, garnet; Tmt, titanomagnetite; Ilm, ilmenite; Rut, rutile; Sph, sphene. Run durations are between 25 h at 1,000 °C to 144 h at 841 °C, see ref. 9 for methods. Glass compositions Val55/4 are from ref. 6. The pegmatite vein 41867 A is from ref. 7.

* H₂O loaded to the capsule.

† Total Fe calculated as FeO.

‡ Dry starting glass composition used in crystallization experiments.

Table 1 Experimental and natural glass/vein chemical compositions and experimental phase proportions

Representative glass (that is, quench melt) compositions are listed in [Table 1](#). In a An-Ab-Or projection ([Fig. 1](#)), the glass compositions define a trend parallel to the albite-anorthite join. Glasses produced at both 10 and 20 kbar cluster in the tonalitic field, whereas those obtained at 30 kbar fall within the trondhjemite field. The latter have Na₂O contents in excess of 7 wt%, increasing as temperature decreases, and are clearly peralkaline (molar Al₂O₃/Na₂O+K₂O+CaO<1), whereas all the glasses produced at lower pressures tend to be either metaluminous or slightly peraluminous. The reason for the large compositional jump observed between 10-20 and 30 kbar is that amphibole breaks down in this pressure range, and the modal abundances of garnet and clinopyroxene increase. The Na₂O released by the amphibole breakdown is not fully balanced by the increase in the Na content of clinopyroxene (that is, the jadeite component) and thus, for a given T and bulk H₂O content, the liquid becomes more sodic and less calcic as pressure increases beyond the amphibole stability field. Extensive garnet crystallization at 30 kbar also lowers the Al₂O₃ content of the glasses relative to those at 20 kbar ([Table 1](#)). In the CaO versus Na₂O diagram ([Fig. 2](#)), all glasses so far produced in H₂O-added melting experiments on basalts up to 30 kbar [11-13](#) display a broad negative trend that completely overlaps that field of natural glasses or veins believed to be slab melts [6,7](#), in contrast to results from dehydration melting experiments [13-15](#) ([Fig. 2b](#)). A similar feature is observed for other elements such as FeO or TiO₂. The melt TiO₂ content ([Table 1](#)) agrees with existing rutile solubility models [16](#). At pressures lower than 10 kbar the glasses are characterized by high CaO and relatively low Na₂O contents, as well as a high Al₂O₃ content, due to the lack of both plagioclase and garnet in the residue [12,13](#). In contrast, glasses obtained in the pressure range 20-30 kbar coexist with an eclogitic residue. Results obtained at 30 kbar match the composition of metasomatic veins the most closely, in particular the extreme Na₂O enrichment. This suggests that these experimental conditions-temperatures below 900°C and bulk water content around 5-6 wt%, thus requiring the presence of a fluid phase before melting-reproduce those of vein production. The origin of free water may lie in the dehydration of serpentinite beneath the oceanic crust of the subducted plate [17,18](#). The water thus produced is expected to dissolve in the melt, except at low melt fraction where there may be enough water to saturate the melt. Clearly, these conditions of slab melting better approach those predicted from thermal modelling [10](#) than do the conditions of the dehydration melting experiments. Therefore, we conclude that free water is needed at depth for slab melting to occur in present-day subduction zones. Experimental constraints on slab melts predict meltwater contents of at least 15 wt% (ref. [9](#)), whereas mass balance calculations on the run products obtained at 30 kbar yield meltwater contents between 24 and 30 wt%. The unusual major-element composition of slab melts appears to be more characteristic of their origin than their trace element signature. In particular, melting in the garnet field of basalt underplated beneath the continental crust may yield silicic melts with trace element patterns similar to those of slab melts [19](#), but they will be granitic instead of trondhjemitic because less water is available in the lower crust and dehydration melting reactions are favoured.

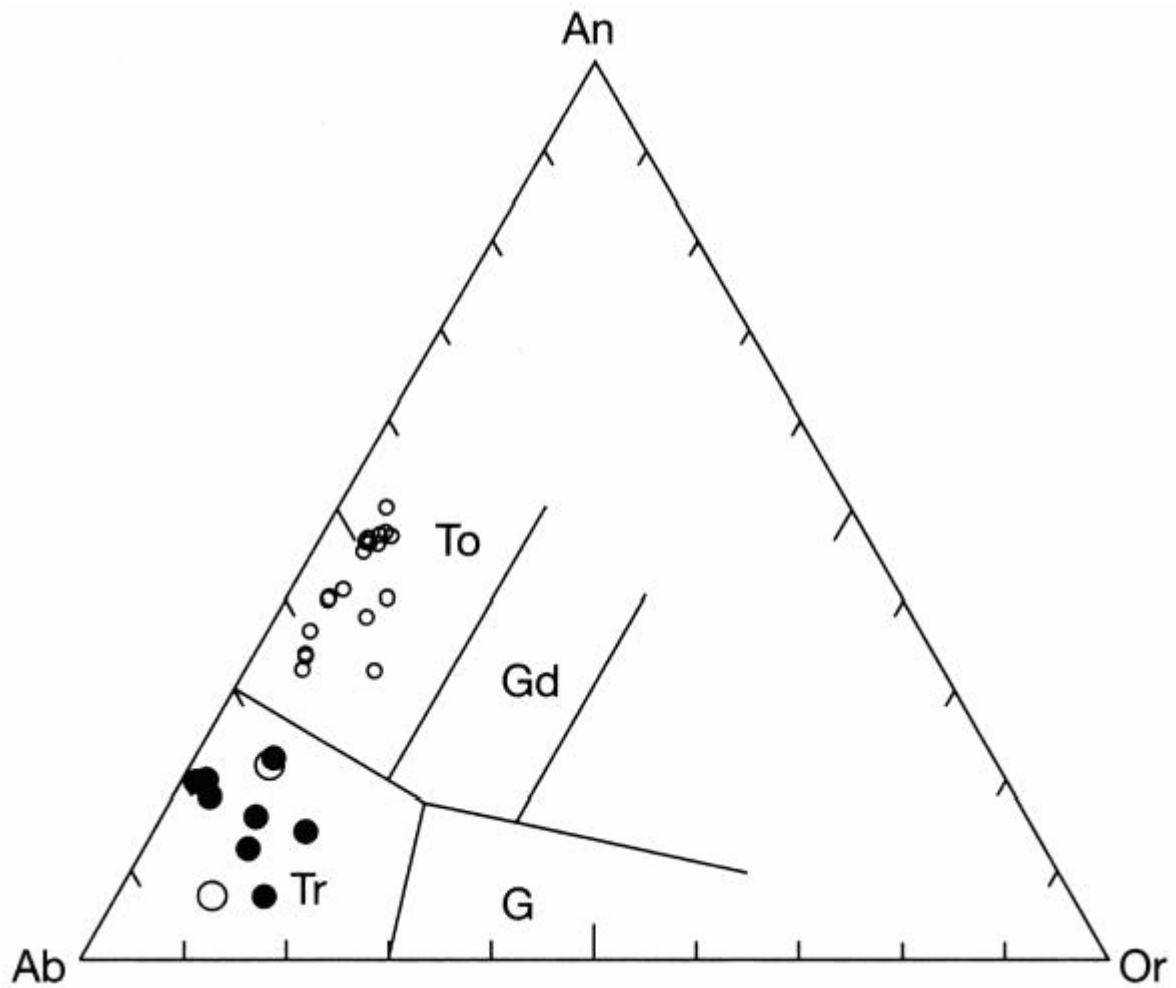


Figure 1 Normative (CIPW) compositions of natural slab melts (black circles) and of glasses obtained in melting experiments of basalts with added H₂O (this study). Small white circles represent experiments done at or below 20 kbar and large white circles experiments at 30 kbar. An, anorthite; Ab, albite; Or, orthoclase; To, tonalite, Tr, trondhjemite; Gd, granodiorite; G, granite.

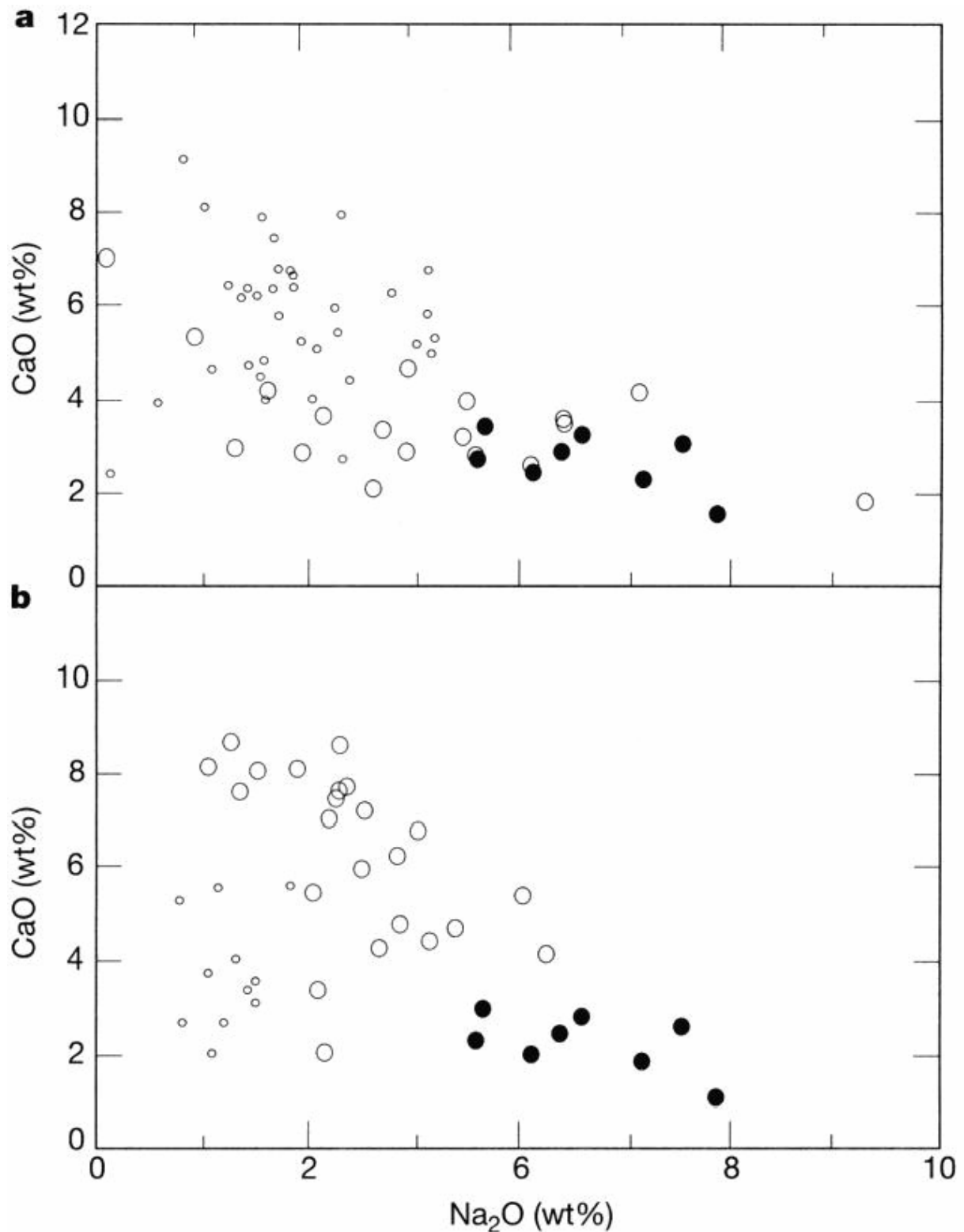


Figure 2 Na₂O versus CaO plot of natural slab melts 6,7 (black circles) and of glasses obtained in melting experiments (large white circles) using N-MORB type basalt protoliths. a, Glasses obtained in melting experiments with added H₂O (this study and refs 11-13). b, Glasses obtained in dehydration melting experiments 13-15. Small white circles represent experiments done at or below 10 kbar (in both a and b).

Interaction of such trondhjemitic melts with the overlying mantle wedge has been experimentally simulated at 15 kbar, as in previous similar works 20-22, by mixing a synthetic equivalent of the trondhjemite melt obtained during melting of the basalt with a forsteritic olivine (Fo 90), the dominant mineral in peridotites (Table 2). Liquid and mineral

compositions are listed in [Table 2](#). At 1,000°C, an orthopyroxene + liquid assemblage is produced. At 900 °C, in addition to orthopyroxene, both an amphibole and phlogopite are reaction products of the interaction process. At both temperatures, olivine appears not to be stable, although residual crystals persist. Glass compositions are notably more magnesian and silica-poor than the starting trondhjemite, being dacitic instead of rhyolitic, yet their strongly alkali-rich character still persists or even increases ([Table 2](#)). An increase of the Mg/(Mg+Fe) ratio of the liquid together with massive orthopyroxene crystallization are key features of geochemical models of slab melt/peridotite interaction [23](#) and reproduce diagnostic petrographic characteristics of slab-derived magmas [24](#). In addition, the compositions of amphibole and phlogopite, the two main minerals diagnostic of mantle metasomatism, are remarkably similar to those of natural amphibole and phlogopite found in some Alpine-type peridotite complexes believed to have been pervasively metasomatized by infiltrating slab-derived melts [8](#). In particular, both minerals are characterized by high Mg/(Fe+Mg) ratios and display strong enrichment in Na₂O. Given the large number of parameters involved in the interaction process (that is, compositions of peridotite and infiltrating melt, solid/melt ratio, pressure and temperature of the interaction) it is expected that metasomatic phases display a wide range of composition, as in previous laboratory studies [20-22](#) and natural occurrences [25](#). However, the sodic character of metasomatic minerals can be related to the sodic character of the metasomatizing agent, as evidenced by the trondhjemite veins of Kamchatka. Variable Na₂O enrichment of metasomatic minerals would trace the more or less trondhjemitic character of the percolating melt which in turn depends on the conditions of its genesis, in particular pressure, as shown above. The similarity between the compositions of experimental and natural phases suggests that the interaction experiments reproduce the conditions of mantle metasomatism and in particular the chemistry of the metasomatic agent.

	SiO ₂	TiO ₂	Al ₂ O ₃	FeO*	MgO	CaO	Na ₂ O	K ₂ O	NiO	total
Starting end members										
Trondh	71.2	0.5	16.9	0.5	0.1	1.5	7.8	1.5	–	100
Olivine	40.8	–	0.1	10.8	47.8	–	–	–	0.3	99.7
Run product compositions										
1,000 °C, 7.8 wt% H ₂ O†										
Gl	65.9	0.5	17.6	1.0	2.9	1.4	8.9	1.6	<0.1	92.2
Opx	58.3	0.1	0.8	3.8	35.3	0.9	0.1	–	0.2	99.5
900 °C, 9.6 wt% H ₂ O										
Gl	64.2	0.5	18.8	1.1	1.5	1.0	11.4	1.5	<0.1	89.6
Opx	57.7	0.1	0.8	6.0	33.4	0.9	0.2	–	0.2	99.4
Phl	41.9	1.6	14.2	2.9	24.4	–	2.7	6.3	0.2	94.4
Amph	50.9	1.7	6.1	3.5	20.6	8.9	4.2	0.3	0.1	96.5
Hydrous phases in metasomatized peridotite‡										
Phl	41.9	0.7	14.3	2.8	25.3	0.0	1.6	6.5	–	94.6
Amph	49.9	0.4	7.5	3.4	20.2	10.2	3.8	0.6	–	98.3

Mixtures consisted in 60 wt% trondhjemite dry glass and 40 wt% olivine. Analyses given in wt%. Glass analyses normalized to 100% anhydrous. Gl, glass; Opx, orthopyroxene; Phl, phlogopite; Amph, amphibole; trondh, trondhjemite.

* Total Fe calculated as FeO.

† Bulk H₂O content of the charge and temperature of the interaction experiment.

‡ Phlogopite and amphibole from the Finero peridotite massif.

Table 2 Glass and mineral compositions from trondhjemite-Fo90 olivine interactions at 15 kbar

The above experimental evidence bears direct constraints on metasomatic processes in arc settings. The low-temperature, fluid-present melting conditions above suggest that slab melting can occur in presently subducting plates, as previously suggested on geochemical grounds [5](#). Although slab melting was originally supposed to have occurred mostly during the Archaean [26](#) or only when young and hot plate subducts in modern settings [5](#), the increasing number of occurrences of magmas with slab melt signatures worldwide [27](#) suggests that conditions for slab melting are realized more often than currently believed. The slab melt contribution as a metasomatizing agent of the sub-arc mantle may thus be important. Mass balance calculations indicate that 15 g of trondhjemite yields 1 g of phlogopite upon reaction with peridotite, whereas 180 g of hydrous fluid are needed to precipitate an equivalent mass

of phlogopite during a peridotite-fluid interaction 28. Thus, as far as metasomatic mineral production is concerned, slab-derived melts are about ten times more efficient than hydrous fluids. Consequently, the geochemical imprint of slab melts can be important in modern arc settings even if slab melting is probably less common than slab dehydration 29. In particular, calculated trace-element concentrations in slab melts in equilibrium with an eclogitic residue (garnet+clinopyroxene+/-rutile) are enriched in light-rare-earth and large-ion-lithophile elements and depleted in high-field-strength elements (La/Nb=7; Ba/Nb=186; ref. 22) and will impart their geochemical signature to magmas generated in the mantle wedge 24.

The overall slab contribution to the source of arc magmas is estimated not to exceed 3% on the basis of oxygen isotope systematics of oceanic arc lavas 30. The reason why slab melts (adakites) are found in volumetrically minor amounts relative to regular calc-alkaline arc magmas may be either physical or chemical. Silica-rich liquids can equilibrate with mantle minerals provided they are alkali-rich 31. This alkali effect is most pronounced at pressures below 20 kbar, however. At 30 kbar, liquids in equilibrium with mantle peridotite are not alkali-rich. The alkali content of experimental slab melts increases with pressure, and comparison with their natural equivalents has shown that slab melting is likely to occur at 30 kbar or more. Thus, besides the physical aspects of magma transport and percolation, any alkali-rich slab melt produced at or beyond 30 kbar, as in our experiments, will be in strong chemical disequilibrium with the overlying mantle wedge. Extensive interaction between these melts and peridotite is expected to occur and the alkali-rich, silica-rich melts will have little opportunity to reach shallower levels unchanged. On this basis, preservation or eruption of genuine slab melts can be anticipated to be difficult, accounting for their relative rarity in most arcs. The chemical differences observed between experimentally produced slab-melts and those erupted or preserved as melt inclusions 32 must in fact reflect various extents of interaction of the latter with mantle material during upward migration.

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